

## Article

# Oscillatory Behaviour of Ni Supported on ZrO<sub>2</sub> in the Catalytic Partial Oxidation of Methane as Determined by Activation Procedure

Daniela Pietrogiacomi <sup>1,\*</sup>, Maria Cristina Campa <sup>2</sup>, Ida Pettiti <sup>1</sup>, Simonetta Tuti <sup>3</sup>, Giulia Luccisano <sup>1</sup>, Leandro Ardeman <sup>1</sup>, Igor Luisetto <sup>4</sup> and Delia Gazzoli <sup>1</sup>

<sup>1</sup> Chemistry Department, Sapienza University of Rome, P.le Aldo Moro 5, 00185 Rome, Italy; ida.pettiti@uniroma1.it (I.P.); giulia.luccisano@uniroma1.it (G.L.); leandrod@yaho.com (L.A.); delia.gazzoli@uniroma1.it (D.G.)

<sup>2</sup> Institute for the Study of Nanostructured Materials (ISMN), National Research Council (CNR), Sapienza University of Rome, P.le Aldo Moro 5, 00185 Rome, Italy; mariacristina.campa@cnr.it

<sup>3</sup> Science Department, Roma Tre University, Via della Vasca Navale 79, 00146 Rome, Italy; simonetta.tuti@uniroma3.it

<sup>4</sup> Italian National Agency for New Technologies, Energy and Sustainable Economic Development (ENEA), Casaccia Research Centre, Via Anguillarese 301, 00123 Rome, Italy;igor.luisetto@enea.it

\* Correspondence: daniela.pietrogiacomi@uniroma1.it; Tel.: +39-06-49913304

**Citation:** Pietrogiacomi, D.; Campa, M.C.; Pettiti, I.; Tuti, S.; Luccisano, G.; Ardeman, L.; Luisetto, I.; Gazzoli, D. Oscillatory Behaviour of Ni Supported on ZrO<sub>2</sub> in the Catalytic Partial Oxidation of Methane as Determined by Activation Procedure. *Materials* **2021**, *14*, 2495. <https://doi.org/10.3390/ma14102495>

Academic Editors: Kai Yan and Barbara Pawelec

Received: 31 March 2021

Accepted: 6 May 2021

Published: 12 May 2021

**Publisher's Note:** MDPI stays neutral with regard to jurisdictional claims in published maps and institutional affiliations.



**Copyright:** © 2021 by the authors. Licensee MDPI, Basel, Switzerland. This article is an open access article distributed under the terms and conditions of the Creative Commons Attribution (CC BY) license (<http://creativecommons.org/licenses/by/4.0/>).

**Abstract:** Ni/ZrO<sub>2</sub> catalysts, active and selective for the catalytic partial oxidation of methane to syngas (CH<sub>4</sub>-CPO), were prepared by the dry impregnation of zirconium oxyhydroxide (Z<sub>hy</sub>) or monoclinic ZrO<sub>2</sub> (Z<sub>m</sub>), calcination at 1173 K and activation by different procedures: oxidation-reduction (*ox-red*) or direct reduction (*red*). The characterization included XRD, FESEM, in situ FTIR and Raman spectroscopies, TPR, and specific surface area measurements. Catalytic activity experiments were carried out in a flow apparatus with a mixture of CH<sub>4</sub>:O<sub>2</sub> = 2:1 in a short contact time. Compared to Z<sub>m</sub>, Z<sub>hy</sub> favoured the formation of smaller NiO particles, implying a higher number of Ni sites strongly interacting with the support. In all the activated Ni/ZrO<sub>2</sub> catalysts, the Ni–ZrO<sub>2</sub> interaction was strong enough to limit Ni aggregation during the catalytic runs. The catalytic activity depended on the activation procedures; the *ox-red* treatment yielded very active and stable catalysts, whereas the *red* treatment yielded catalysts with oscillating activity, ascribed to the formation of Ni<sup>δ+</sup> carbide-like species. The results suggested that Ni dispersion was not the main factor affecting the activity, and that active sites for CH<sub>4</sub>-CPO could be Ni species at the boundary of the metal particles in a specific configuration and nuclearity.

**Keywords:** supported Ni/ZrO<sub>2</sub> catalysts; Ni metal particles; methane partial oxidation; activation procedure; heterogeneous catalysis

## 1. Introduction

The global energy demand is expected to rise by 30% between today and 2040. In this scenario, the use of natural gas as an energy source is expected to increase by 45% in the next 20 years, according to the World Energy Outlook 2017—International Energy Agency [1]. The conversion of natural gas, mainly constituted by methane, into highly valuable products has become challenging [1–4]. The catalytic conversion of methane to syngas is an attractive research area as syngas is a building block for valuable chemicals and liquid fuels [5–7]. The currently adopted large-scale process to produce syngas is CH<sub>4</sub> steam reforming [8,9], which is energetically expensive due to its endothermic characteristics. Conversely, the catalytic partial oxidation of methane (CH<sub>4</sub>-CPO) is a valid alternative because it is mildly exothermic and can produce syngas in a ratio (H<sub>2</sub>/CO = 2) suitable for methanol or Fischer–Tropsch synthesis [5–7,10–13].

Catalytic systems based on noble (Rh, Pt and Ir) or non-noble metals (Ni, Co and Fe) supported on various oxides have been studied for the CH<sub>4</sub>-CPO reaction [4,5,11,14]. By consensus, nickel supported on oxides is considered a promising catalyst owing to the high methane conversion, high CO and H<sub>2</sub> selectivities and low cost [11,14–20], but it suffers from deactivation. Many studies have shown that deactivation arises from the sintering of Ni particles and/or coke deposition, both depending on metal particle and support features [11,14,21–24]. The choice of support is fundamental, as the metal–support interaction can favour the formation of peculiar active sites at the metal–support boundaries [25,26].

Zirconium oxide is widely used as a support because of its high thermal stability and characteristic textural properties that can be tailored according to different preparation methods and thermal treatments, yielding the formation of monoclinic, metastable tetragonal or mixed tetragonal-monoclinic ZrO<sub>2</sub> phases [27–29]. Many papers have extensively investigated Ni-supported ZrO<sub>2</sub> systems for syngas production focussing on the support preparation methods and on the deposition of Ni precursors to obtain metal particle sizes suitable for high catalytic performances and low carbon deposition [11,14,30–36]. By contrast, the effect of ZrO<sub>2</sub> modifications on the metal dispersion and on the structure–activity relationship has not been clearly established thus far. Studies on the interaction between Ni clusters and ZrO<sub>2</sub> polymorphs have reported that monoclinic ZrO<sub>2</sub> performs better than tetragonal and cubic forms both for cluster dispersion and aggregation inhibition [37,38]. Furthermore, the effect of the activation procedures of catalyst precursors on the Ni metal particle features has not been deeply analysed, although it is generally recognised to affect the strength of the metal–support interaction [39].

In this paper, we studied the catalytic performances for the CH<sub>4</sub>-CPO of nickel-supported systems in which the support was ZrO<sub>2</sub> in the monoclinic modification. The thermal stability of the monoclinic phase up to about 1400 K, able to prevent changes induced by the occurrence of hot spots during CH<sub>4</sub>-CPO reaction, and the chemical inertness between NiO and ZrO<sub>2</sub> are expected to assure catalyst stability. Ni/ZrO<sub>2</sub> catalysts were prepared by the impregnation of two different starting materials (a high-surface-area zirconium oxyhydroxide, Z<sub>hy</sub>, and a low-surface-area monoclinic zirconium oxide, Z<sub>m</sub>) and were activated following two different procedures (oxidation–reduction treatment, *ox-red*, or direct reduction treatment, *red*). The objective of this study was to clarify the effect of the starting material and of the reductive activation procedure on the nature of the active phase and on the catalytic behaviour. To this aim, we compared the catalytic behaviours and features of Ni/ZrO<sub>2</sub> catalysts with those of bare monoclinic ZrO<sub>2</sub> and of an unsupported Ni powder sample. Materials were characterized by various techniques, including atomic absorption spectroscopy (AAS), X-ray diffraction (XRD), specific surface area (BET method) and pore volume analysis, field-emission scanning electron microscopy (FESEM), temperature-programmed reduction (TPR), and *in situ* transmission Fourier-transform infrared (FTIR) and Raman spectroscopies. The correlation between the characterization and catalytic activity results allowed some hypotheses to be formulated on the active sites for the CH<sub>4</sub>-CPO and on the consequent oscillating catalytic behaviour.

## 2. Materials and Methods

### 2.1. Materials

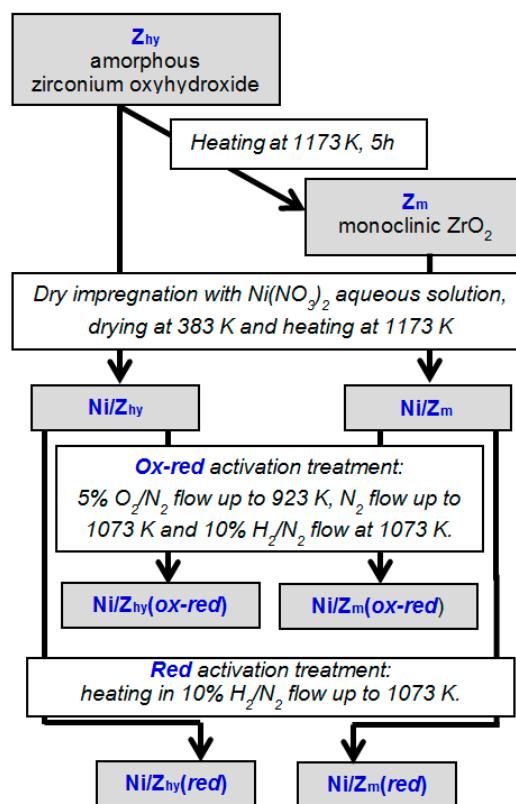
Zirconium oxyhydroxide (Z<sub>hy</sub>) was prepared via precipitation from a ZrOCl<sub>2</sub> solution with ammonia. After separation, the solid was washed with water until a Cl<sup>−</sup>-negative test (no opalescence in the liquid after addition of AgNO<sub>3</sub>). The test was conducted on the washing water and on small portions of solid removed from the batch and dissolved in dilute HNO<sub>3</sub>. After washing, the solid was dried at 383 K for 24 h. A

portion of this material was calcined at 1173 K (heating rate:  $0.25\text{ K}\cdot\text{min}^{-1}$ ) for 5 h, indicated as  $Z_m$  ( $m$  stands for monoclinic).

$\text{NiO}/\text{ZrO}_2$  catalyst precursors (about 2 or 5 Ni wt.%) were prepared by the dry impregnation of  $Z_{hy}$  or  $Z_m$  material with an  $\text{Ni}(\text{NO}_3)_2$  aqueous solution, and they were dried at 383 K and subsequently calcined at 1173 K for 5 h; the heating rate was  $0.25\text{ K}\cdot\text{min}^{-1}$  for the  $Z_{hy}$ -based samples and was  $5.0\text{ K}\cdot\text{min}^{-1}$  for those based on  $Z_m$ . The adoption of such a slow heating ramp ( $0.25\text{ K}\cdot\text{min}^{-1}$ ) for the  $Z_{hy}$ -based samples was necessary to obtain crystalline  $\text{ZrO}_2$  with the highest possible surface area, avoiding a marked local increase in temperature due to the exothermicity of the crystallization process.  $\text{NiO}/\text{ZrO}_2$  catalyst precursors were stored in plastic vials at room temperature in air and were labelled as  $x\text{NiO}/Z_a$ , where  $x$  indicates the Ni content (Ni wt.%) and  $Z_a$  specifies the starting material used for the impregnation,  $Z_{hy}$  or  $Z_m$ .

$\text{Ni}/\text{ZrO}_2$  catalysts were obtained by the in situ reduction of precursors following different activation procedures: (i) oxidation in a 5%  $\text{O}_2/\text{N}_2$  flow up to 923 K for 1 h, purging with  $\text{N}_2$  up to 1073 K, and reduction in a 10%  $\text{H}_2/\text{N}_2$  flow at 1073 K for 1 h (*ox-red* activation); (ii) direct reduction in a 10%  $\text{H}_2/\text{N}_2$  flow up to 1073 K for 1 h (*red* activation). Samples were named  $x\text{Ni}/Z_a$  (*ox-red*) and  $x\text{Ni}/Z_a$  (*red*), respectively. The catalyst preparation procedures are sketched in Scheme 1.

Polycrystalline nickel oxide was obtained by the slow decomposition of  $\text{Ni}(\text{CH}_3\text{COO})_2$  in air at 1173 K (heating rate:  $0.25\text{ K}\cdot\text{min}^{-1}$ ) for 5 h.



**Scheme 1.** Preparation procedures of  $\text{Ni}/\text{ZrO}_2$  catalysts.

## 2.2. Characterization

The Ni content of catalysts was determined by AAS (SpectrAA 220, Varian Australia Pty Ltd., Mulgrave, Australia). X-ray diffraction (XRD) patterns were obtained with a Philips PW 1729 diffractometer (Malvern Panalytical Ltd., Malvern, UK) using  $\text{Cu K}\alpha$  (Ni-filtered) radiation in a  $2\theta$  range of  $10\text{--}70^\circ$  (step size:  $0.02^\circ$ ; time per step: 1.25 s). The mean crystallite diameter ( $d$ ) was calculated by the Scherrer equation [40] as  $d = K\lambda/\beta\cos\theta$ , where  $K$  is a shape constant equal to 0.9,  $\lambda$  is the X-ray wavelength used, and  $\beta$

is the effective linewidth (FWHM) of the observed X-ray reflection, obtained by a curve-fitting procedure after Warren's correction for instrumental broadening and background subtraction.

Specific surface area (BET method) and textural properties were determined by the adsorption/desorption of N<sub>2</sub> at 77 K using a Micromeritics ASAP 2010 Analyzer (Norcross, GA, USA) after sample outgassing at 473 K for 2 h via thermally controlled heating mantles, up to a residual pressure lower than 0.8 Pa. The pore distribution was determined by the BJH method [41] from the adsorption isotherm. The total pore volume was obtained by the rule of Gurvitsch [42]. The uncertainty was  $\pm 0.5 \text{ m}^2 \cdot \text{g}^{-1}$  for the specific surface area values and  $\pm 0.005 \text{ cm}^3 \cdot \text{g}^{-1}$  for the total pore volume values.

Field-emission scanning electron microscopy (FESEM) images were obtained by using an AURIGA Zeiss 405 HR-FESEM instrument (Oberkochen, Baden-Württemberg, Germany), equipped with an energy-dispersive X-ray spectroscopy (EDXS) Bruker apparatus for elemental detection, which is able to discriminate particles with a minimum diameter of about 5 nm.

To evaluate the Ni dispersion, D (percent ratio of the exposed Ni atoms to the total Ni atoms), FESEM images were processed using the free ImageJ software version 1.53h (National Institutes of Health, Bethesda, MD, USA) [43] to measure the particle diameters, d<sub>i</sub>. The particle diameter distribution was established by measuring all the particles observed in at least 5 images of each sample. The error in the particle size measurement due to the pixel dimensions of FESEM images (depending on the resolution at different magnifications) was  $\leq 10\%$ . The metal dispersion was derived according to [44] using the approximation of spherical particles, taking into account the relationship between D and the mean particle diameter,  $D = 100 \cdot 6 \cdot (v_{Ni}/a_{Ni})/d_{VA}$ , where  $d_{VA} = \sum_{i=1}^n d_i^3 / \sum_{i=1}^n d_i^2$  is the volume-area mean diameter,  $v_{Ni}$  is the volume of a Ni atom in bulk metal ( $0.01095 \text{ nm}^3$ ), and  $a_{Ni}$  is the area occupied by a surface Ni atom ( $0.0651 \text{ nm}^2$ ).

Temperature-programmed reduction (TPR) experiments were performed using a Thermo Scientific TPDRO1100 instrument (Waltham, MA, USA). The H<sub>2</sub>-TPR analysis was conducted by flowing a 5% H<sub>2</sub>/Ar mixture ( $10 \text{ cm}^3 \cdot \text{min}^{-1}$ ) through the sample (100 mg) from 323 to 1073 K (heating rate  $10 \text{ K} \cdot \text{min}^{-1}$ ) and keeping the sample at 1073 K for 1 h. The H<sub>2</sub> consumption was measured by a TCD detector, calibrated by the reduction of a known amount of CuO (99.99% purity; Sigma Aldrich, Saint Louis, MO, USA). Before flowing into the TCD, the H<sub>2</sub>O produced during the reduction was removed by a soda lime trap.

Raman spectra were recorded at room temperature, in back-scattering geometry, with an inVia Renishaw micro-Raman spectrometer (Wotton-under-Edge, Gloucestershire, UK), using the 488.0 nm emission line from an Ar ion laser as the exciting source. The power of the incident beam was about 5 mW. Repeated accumulations (10 or 20 scans  $\times$  10 s) were generally acquired on at least four regions of each sample using 20 $\times$  or 5 $\times$  objectives to check the sample homogeneity. Spectra were calibrated using the 520.5 cm<sup>-1</sup> line of a silicon wafer. Spectra processing included baseline removal and curve fitting using a Gauss–Lorentz cross-product function by Peakfit 4.12 (Systat Software Inc., San Jose, CA, USA, 2007).

Transmission FTIR spectra were recorded with a Perkin Elmer Frontier spectrometer (Milano, Italy), equipped with a MCT detector operating with a resolution of 4 cm<sup>-1</sup>. The powdered sample, crushed and pelleted (pressure:  $1.5 \times 10^4 \text{ kg} \cdot \text{cm}^{-2}$ ) in a self-supporting wafer of about  $15 \text{ mg} \cdot \text{cm}^{-2}$ , was inserted in a stainless-steel reactor equipped with CaF<sub>2</sub> windows. The reactor, connected to a flow apparatus, allowed spectra to be recorded during thermal treatments up to 773 K in an oxidative (5% O<sub>2</sub>/N<sub>2</sub>) or reductive (10% H<sub>2</sub>/N<sub>2</sub>) flow to simulate the catalytic activation treatment. At a given temperature, surface species resulting from the various treatments were determined by subtracting reference spectra from those recorded after specific treatments. The reference spectra were collected in flowing helium, on samples heated up to 773 K and cooled down to the desired temperature (298–773 K).

### 2.3. Catalytic Activity

The CH<sub>4</sub>-CPO reaction was studied in a flow apparatus at atmospheric pressure. The feeding section included independent mass flow controller meters (MKS mod. 1259, driven by a four-channel unit MKS mod. 647C) and a glass ampoule for gas mixing before entering the reactor. The fixed-bed tubular reactor was made of two coaxial quartz tubes (i.d. 20 and 10 mm) to allow preheating of the feed gas. The temperature of the catalytic bed was monitored by a K-type thermocouple, located in a quartz tube (i.d. 4 mm) concentric to the reactor. High-purity gas mixtures (CH<sub>4</sub>/N<sub>2</sub>, O<sub>2</sub>/N<sub>2</sub> and H<sub>2</sub>/N<sub>2</sub> from Rivoira gas, Milano, Italy), and pure N<sub>2</sub> (SOL) were used without further purification. Reactants and products were analysed by a Varian Micro-GC CP-4900 gas chromatograph (Agilent, Santa Clara, CA, USA) equipped with two columns (10 m Molsieve 5A BF, for H<sub>2</sub>, O<sub>2</sub> and CO; 10 m Poraplot Q, for CH<sub>4</sub>, CO<sub>2</sub> and H<sub>2</sub>O) and TCD detectors. Sample masses of 50 mg for ZrO<sub>2</sub> and NiO/ZrO<sub>2</sub> and 3.0 mg for unsupported NiO (nickel amount comparable to that contained in 50 mg of a catalyst with 5 wt.% Ni) were deposited on a ceramic wool bed in the reactor. The height of the catalytic bed in the reactor was ≤1 mm. Before catalytic measurements of Z<sub>m</sub>, unsupported NiO and NiO/ZrO<sub>2</sub> catalysts were submitted to the in situ reduction procedure, *ox-red* or *red*, as specified in the Materials section (Section 2.1.).

Typical catalytic runs consisted of steady-state measurements with a mixture of 2% CH<sub>4</sub>, 1% O<sub>2</sub>, and N<sub>2</sub> as balance, in the temperature range of 1023–723 K, leaving the catalysts in a continuous flow of the reactant mixture. The temperature was changed in a random sequence, maintaining a constant temperature for about 15 min (three consecutive Micro-GC analyses). To check the activity reproducibility, some temperature values were explored two or more times in the same run. For each catalyst, up to four runs were carried out by applying an *ox-red* activation treatment between two consecutive runs. The stability of the activity was tested in ad hoc runs following the reaction at a set temperature as a function of time on stream up to 5 h. The total flow rate was 150 cm<sup>3</sup> (STP) min<sup>-1</sup> (space velocity: 1.8·10<sup>5</sup> NL·kg<sub>cat</sub><sup>-1</sup>·h<sup>-1</sup>; contact time  $\tau = 9$  ms). At constant temperature, the linear increase in conversions at increasing contact time ( $\tau$ ) indicated that, under our experimental conditions, the reaction was under kinetic control without external diffusion effects (Figure S1 in the Supplementary Material). All experiments yielded a satisfactory carbon, hydrogen and oxygen balance (100 ± 5%). Samples after catalytic activity tests were named *used* catalysts.

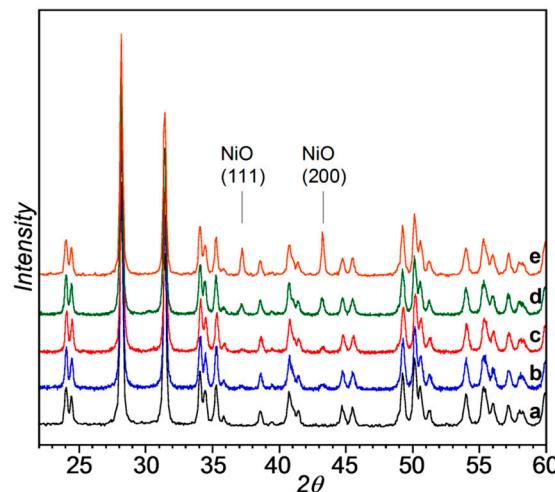
The catalytic activity and selectivity for the CH<sub>4</sub>-CPO reaction (CH<sub>4</sub> + 1/2O<sub>2</sub> → CO + 2H<sub>2</sub>) were discussed considering the simultaneous occurrence of methane total combustion (CH<sub>4</sub> + 2O<sub>2</sub> → CO<sub>2</sub> + 2H<sub>2</sub>O), and the water gas shift (WGS) reaction (CO + H<sub>2</sub>O → CO<sub>2</sub> + H<sub>2</sub>). The percent CH<sub>4</sub> conversion was calculated as 100 (CH<sub>4</sub> molecules consumed)/(CH<sub>4</sub> molecules injected). The percent H<sub>2</sub> yield was calculated as 100 (H<sub>2</sub> molecules produced)/(1/2 CH<sub>4</sub> molecules injected), and CO or CO<sub>2</sub> yields were calculated as 100 (CO or CO<sub>2</sub> molecules produced)/(CH<sub>4</sub> molecules injected). The H<sub>2</sub> selectivity was calculated as 100 (H<sub>2</sub> molecules produced)/(1/2 CH<sub>4</sub> molecules converted). The CO selectivity was calculated as 100 (CO molecules produced)/(CH<sub>4</sub> molecules converted). The rate of H<sub>2</sub> production (R<sub>H<sub>2</sub></sub>/molecules s<sup>-1</sup>·g<sup>-1</sup>) was calculated from H<sub>2</sub> molecules produced in experiments in which the conversion never exceeded 30%. The rate was correlated to the number of exposed Ni atoms per gram of catalyst (N<sub>Ni(exp)</sub>), obtained according to the expression  $N_{Ni(exp)} = (D \cdot w_{Ni} \cdot N_A) / (10,000 \cdot M_{Ni})$ , where D is the Ni dispersion value, w<sub>Ni</sub> is the Ni loading (wt.%), N<sub>A</sub> is Avogadro's number and M<sub>Ni</sub> is the Ni molar mass (58.69 g·mol<sup>-1</sup>).

### 3. Results and Discussion

#### 3.1. Characterization

##### 3.1.1. Zirconium Oxide Support

Amorphous zirconium oxyhydroxide,  $Z_{hy}$ , heated at 1173 K transformed into monoclinic  $ZrO_2$ ,  $Z_m$ , as shown by XRD (JCPDF card 37-1484) (Figure 1a). The  $Z_m$  average crystallite size determined by the Scherrer equation [41] was 49 nm; the (−111) reflection at 28.4° was considered for the fitting procedure. Specific surface area measurements and textural analysis showed that  $Z_{hy}$  was a micro-mesoporous material, with a specific surface area of  $360 \text{ m}^2 \cdot \text{g}^{-1}$ , and that monoclinic  $ZrO_2$ ,  $Z_m$ , was a meso-macroporous material, with a specific surface area of  $5.6 \text{ m}^2 \cdot \text{g}^{-1}$  (Table 1).



**Figure 1.** XRD patterns of monoclinic  $ZrO_2$  and  $\text{NiO}/\text{ZrO}_2$  catalyst precursors. (a)  $Z_m$ ; (b) 1.8NiO/ $Z_m$ ; (c) 1.7NiO/ $Z_{hy}$ ; (d) 4.8NiO/ $Z_{hy}$ ; (e) 5.1NiO/ $Z_m$ .

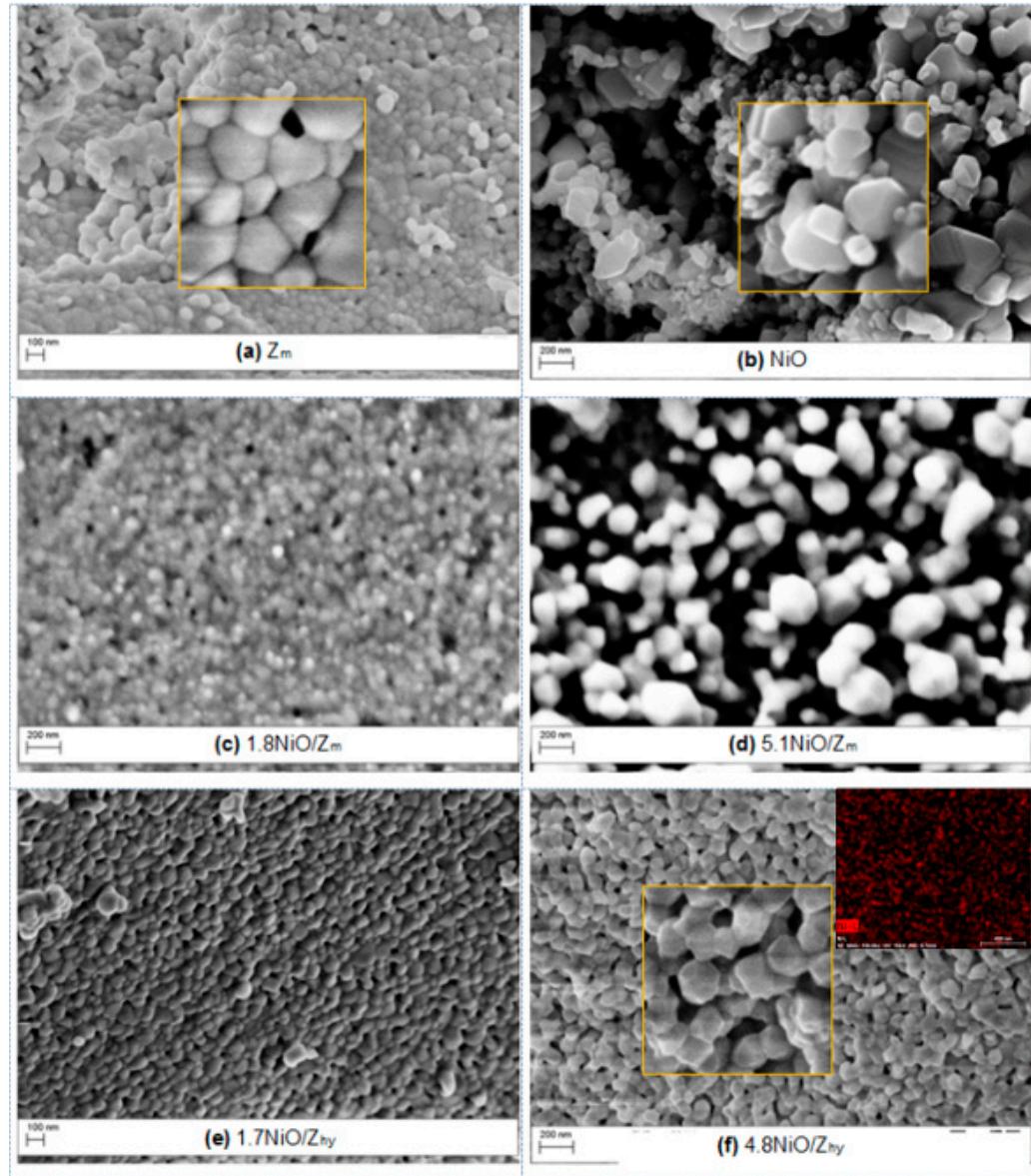
**Table 1.** Structural and textural features of precursors and catalysts (activated or used): specific surface area (S.A.), total pore volume ( $V_{tot}$ ) and crystallite size ( $d_{XRD}$ ) of  $ZrO_2$ ,  $\text{NiO}$  and  $\text{Ni}$ .

Precursors	Catalysts	S.A. <sup>1</sup> ( $\text{m}^2 \cdot \text{g}^{-1}$ )	$V_{tot}$ <sup>2</sup> ( $\text{cc} \cdot \text{g}^{-1}$ )	Crystallite Size, $d_{XRD}$		
				$\text{ZrO}_2$	$\text{NiO}$	$\text{Ni}$ act <sup>3</sup> Used
$Z_{hy}$	-	360	0.33	-	-	-
-	$Z_m$	5.6	0.041	49	-	-
NiO (unsupported)	-	-	-	-	70	-
-	Ni (unsupported)	-	-	-	-	70
1.8NiO/ $Z_m$	-	2.7	0.024	-	22	-
-	1.8Ni/ $Z_m$ (ox-red)	-	-	-	-	18
-	1.8Ni/ $Z_m$ (red)	-	-	-	-	25
5.1NiO/ $Z_m$	-	3.3	0.031	50	36	-
-	5.1Ni/ $Z_m$ (ox-red)	-	-	-	-	30
-	5.1Ni/ $Z_m$ (red)	-	-	-	-	35
1.7NiO/ $Z_{hy}$	-	5.4	0.040	-	32	-
-	1.7Ni/ $Z_{hy}$ (ox-red)	-	-	-	-	20
-	1.7Ni/ $Z_{hy}$ (red)	-	-	-	-	25
4.8NiO/ $Z_{hy}$	-	4.4	0.044	52	26	-
-	4.8Ni/ $Z_{hy}$ (ox-red)	-	-	-	-	23

	4.8Ni/Z <sub>hy</sub> (red)	-	-	-	-	-	24	30
--	-----------------------------	---	---	---	---	---	----	----

<sup>1</sup>  $\pm 0.5 \text{ m}^2 \cdot \text{g}^{-1}$ ; <sup>2</sup>  $\pm 0.005 \text{ cm}^3 \cdot \text{g}^{-1}$ ; <sup>3</sup> act, activated catalysts.

FESEM micrographs of the Z<sub>m</sub> support revealed ZrO<sub>2</sub> particles having sizes in the range of 30–100 nm (Figure 2a and Table 2).



**Figure 2.** FESEM images of monoclinic ZrO<sub>2</sub>, unsupported NiO and NiO/ZrO<sub>2</sub> catalyst precursors. Magnifications of selected areas are shown in the insets of (a), (b) and (f). The inset with red dots in (f) refers to the Ni EDX elemental map.

**Table 2.** FESEM images analysis for  $Z_m$ , NiO and Ni.

Precursors	Catalysts	Particle Size, $d_{FESEM}$ (nm)			Ni Particle Number $\Sigma n_i$ <sup>1</sup>	Ni Particle Mean Diameter		Ni Dispersion D <sup>4</sup> (%)	Exposed Ni Atoms N <sub>Ni(exp)</sub> <sup>5</sup> (atoms/g)
		$ZrO_2$	NiO	Ni Activated Used		$d_{LN}$ <sup>2</sup> (nm)	$d_{VA}$ <sup>3</sup> (nm)		
-	$Z_m$	30–100	-	-	-	-	-	-	-
NiO (unsupported)	-	-	20–400	-	-	-	-	-	-
-	Ni (unsupported)			40–1500 50–2000	-	-	-	-	-
1.8NiO/ $Z_m$	-	-	40–60	-	-	-	-	-	-
-	1.8Ni/ $Z_m$ (ox-red)	-	-	5–75 20–80	1043	$36 \pm 15$	48	2.1	$3.9 \cdot 10^{18}$
-	1.8Ni/ $Z_m$ (red)	-	-	20–75 20–80	-	-	-	-	-
5.1NiO/ $Z_m$	-	-	80–300	-	-	-	-	-	-
-	5.1Ni/ $Z_m$ (ox-red)	-	-	30–1000 50–700	502	$2.6 \cdot 10^2$	$5.5 \cdot 10^2$	0.18	$9.4 \cdot 10^{17}$
-	5.1Ni/ $Z_m$ (red)	-	-	30–700 35–700	-	-	-	-	-
1.7NiO/ $Z_{hy}$	-	-	n.d. <sup>6</sup>	-	-	-	-	-	-
-	1.7Ni/ $Z_{hy}$ (ox-red)	-	-	10–35 15–25	25	$18 \pm 7$	22	4.6	$7.1 \cdot 10^{18}$
-	1.7Ni/ $Z_{hy}$ (red)	-	-	5–30 15–20	-	-	-	-	-
4.8NiO/ $Z_{hy}$	-	-	n.d. <sup>6</sup>	-	-	-	-	-	-
-	4.8Ni/ $Z_{hy}$ (ox-red)	-	-	10–45 15–40	188	$24 \pm 7$	28	3.6	$1.8 \cdot 10^{19}$
-	4.8Ni/ $Z_{hy}$ (red)	-	-	8–35 15–40	-	-	-	-	-

<sup>1</sup>  $n_i$ , number of metal particles with a specific diameter  $d_i$ ; <sup>2</sup> length-number mean diameter,  $d_{LN} = \Sigma n_i \cdot d_i / \Sigma n_i$  and the corresponding standard deviation  $\sigma$ ; <sup>3</sup> volume-area mean diameter,  $d_{VA} = \Sigma n_i \cdot d_i^3 / \Sigma n_i \cdot d_i^2$ ; <sup>4</sup> Ni dispersion,  $D = 100 \cdot 6 \cdot (v_{Ni}/a_{Ni})/d_{VA}$ ; <sup>5</sup> exposed Ni atoms,  $N_{Ni(exp)} = (D \cdot w_{Ni} \cdot N_A) / (10,000 \cdot M_{Ni})$ ; <sup>6</sup> n.d., not detected.

### 3.1.2. NiO/ZrO<sub>2</sub> Catalyst Precursors

All the NiO/ZrO<sub>2</sub> samples (NiO/ $Z_{hy}$  and NiO/ $Z_m$ ) heated at 1173 K, in addition to the reflection of monoclinic ZrO<sub>2</sub>, showed peaks at 2θ values of 37.3° and 43.3°, corresponding to the (111) and (200) crystallographic planes of cubic NiO (JCPDF card 4-0835), (Figure 1, patterns b–e). For all the NiO/ZrO<sub>2</sub> samples, the NiO average crystallite sizes, determined by the (200) reflection (Scherrer equation [41]), ranged from 22 to 36 nm and was smaller than those of unsupported NiO (70 nm, Table 1), suggesting that an interaction between NiO and the support positively affected the crystallite sizes.

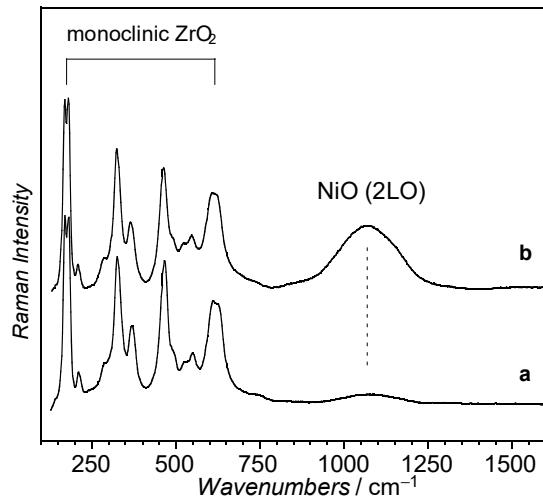
FESEM micrographs of unsupported NiO mainly revealed octahedral truncated particles (inset in Figure 2b) with sizes in the range of 20–400 nm (Figure 2b and Table 2). The images of the NiO/ $Z_m$  samples revealed NiO particles (EDX analysis) with sizes within the range exhibited by unsupported NiO, about 40–60 nm for the 1.8NiO/ $Z_m$  sample and 80–300 nm for the 5.1NiO/ $Z_m$  sample (Figure 2c,d; Table 2). In particular, in the latter sample, NiO-supported particles showed truncated octahedral shapes similar to those of unsupported NiO particles. The NiO particles imaged by FESEM were larger than those determined by XRD, indicating that, on the zirconia surface, NiO was in the form of crystallite aggregates.

The nitrogen adsorption/desorption Type IV isotherms (Figure S2a) showed H1 hysteresis loops characteristic of adsorbents with a narrow distribution of uniform pores. The surface area and total pore volume values of NiO/ $Z_m$  samples were lower than those of the  $Z_m$  support (Table 1), due to the decrease in the meso- and macroporosity, sug-

gesting that some of the NiO particles partially occluded the pores. Accordingly, pore size distribution plots exhibited a mesoporous texture more abundant for  $Z_m$  than for the NiO/ZrO<sub>2</sub> catalyst precursors (Figure S2b).

In the NiO/Z<sub>hy</sub> samples, although EDX analysis detected a homogeneous distribution of Ni species, and NiO crystallite sizes estimated by XRD were above the size detection limit of FESEM (5 nm), no NiO particles were imaged (Figure 2e,f). This fact could be due to (i) the low contrast between supported NiO particles and the zirconia support, and/or (ii) the location of the NiO particles in the ZrO<sub>2</sub> interparticle voids, formed during the calcination of Z<sub>hy</sub> impregnated with Ni(NO<sub>3</sub>)<sub>2</sub>. The similarity of surface area and total pore volume values between NiO/Z<sub>hy</sub> samples and the  $Z_m$  support (Table 1) suggests that, for NiO/Z<sub>hy</sub> samples, pore occlusion of ZrO<sub>2</sub> by NiO did not occur, supporting the hypothesis of small NiO particles. A closer inspection of the 4.8NiO/Z<sub>hy</sub> FESEM image (inset in Figure 2f) indicated particles with truncated octahedral morphologies, unlike the roundish particles of the  $Z_m$  support but similar to those of unsupported NiO particles. This result suggests that, during the calcination treatment, the NiO phase spread as a raft-like layer on the Z<sub>hy</sub> support. This intimate contact between the two oxide phases could account for the low contrast in the FESEM images [45].

Raman spectra confirmed the presence of NiO both in the NiO/ $Z_m$  and NiO/Z<sub>hy</sub> samples (Figure 3). NiO showed Raman bands at about 420 and 560 cm<sup>-1</sup> (one-phonon, 1P: TO and LO modes, respectively) and at about 730, 906 and 1090 cm<sup>-1</sup> (two-phonon, 2P: 2TO, TO + LO and 2LO modes, respectively) [46]. With the decrease in particle size, the 1P bands became more pronounced due to defects, whereas the 2P bands broadened and decreased in intensity [47]. For both 5.1NiO/ $Z_m$  and 4.8NiO/Z<sub>hy</sub> samples, while the 1P and 2P modes in the range of 200–750 cm<sup>-1</sup> were completely obscured by those of monoclinic ZrO<sub>2</sub>, the broad band at about 1090 cm<sup>-1</sup> (2LO mode) clearly indicated the presence of the NiO phase (Figure 3). The much higher intensity of the 2LO band in the 5.1NiO/ $Z_m$  sample than in the 4.8NiO/Z<sub>hy</sub> sample was consistent with the presence of larger NiO particles in the former than in the latter.

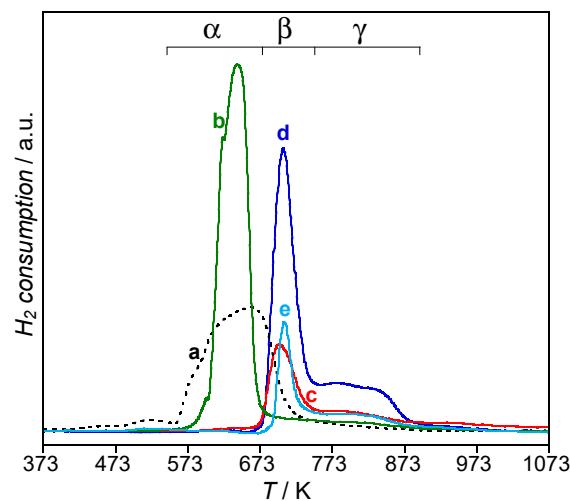


**Figure 3.** Raman spectra of catalyst precursors. (a) 4.8NiO/Z<sub>hy</sub>; (b) 5.1NiO/ $Z_m$ .

Temperature-programmed reduction (TPR) was applied to study the reducibility of NiO species, which is expected to depend on the strength of interaction with the support. In all the samples, NiO species were completely reduced to Ni metal in the temperature range of 473–923 K (total H<sub>2</sub> consumption corresponding to about 2 e/Ni, Table 3).

The reduction profile of the 5.1NiO/ $Z_m$  sample (Figure 4b) showed an intense H<sub>2</sub> consumption peak in the 573–673 K range with a maximum at 640 K, close in position to that of unsupported NiO (Figure 4a), and a very weak H<sub>2</sub> consumption spread up to 923 K. The 1.8NiO/ $Z_m$ , 4.8NiO/Z<sub>hy</sub> and 1.7NiO/Z<sub>hy</sub> samples (Figures 4, profiles c–e) showed a

very similar reduction profile with an intense and narrow peak at about 700 K and a broad hydrogen consumption in the 750–923 K range, with the integrated intensity of the peaks being proportional to the nickel content. Several authors have suggested that for NiO supported on ZrO<sub>2</sub> or modified ZrO<sub>2</sub>, the stronger the interaction, the higher the reduction temperature [32,33,48–51]. Accordingly, the intense peak with a maximum at 640 K was attributed to NiO species not or weakly interacting with the ZrO<sub>2</sub> surface, named  $\alpha$  species; the peak and the envelope in the range of 673–873 K were ascribed to species interacting with the support, named  $\beta$  and  $\gamma$ , with the  $\gamma$  species being more strongly interacting than the  $\beta$  species are [51]. The similarity in shape and position of the reduction profiles of 1.8NiO/Z<sub>m</sub>, 4.8NiO/Z<sub>hy</sub> and 1.7NiO/Z<sub>hy</sub> precursors indicated that the strength of the NiO interaction with the support did not depend on the specific starting material adopted for the impregnation, Z<sub>m</sub> or Z<sub>hy</sub>. As shown in Table 3, for these samples, the H<sub>2</sub> consumption was only due to  $\beta$  and  $\gamma$  NiO species and was much higher than that due to  $\beta$  and  $\gamma$  species in the 5.1NiO/Z<sub>m</sub> sample (100% vs 11%). The lower  $\beta$  and  $\gamma$  species amounts in 5.1NiO/Z<sub>m</sub> than in 1.8NiO/Z<sub>m</sub> precursors (Figure 4, profiles b and c, respectively) indicated that the amount of NiO interacting with the monoclinic ZrO<sub>2</sub> decreased with nickel content, due to saturation of the low number of surface sites available on Z<sub>m</sub> to anchor Ni<sup>2+</sup> species.



**Figure 4.** TPR profiles for NiO and NiO/ZrO<sub>2</sub> catalyst precursors. (a) Unsupported NiO; (b) 5.1NiO/Z<sub>m</sub>; (c) 1.8NiO/Z<sub>m</sub>; (d) 4.8NiO/Z<sub>hy</sub>; (e) 1.7NiO/Z<sub>hy</sub>.  $\alpha$ ,  $\beta$  and  $\gamma$ : temperature ranges for the reduction of NiO species interacting differently with the support.

**Table 3.** H<sub>2</sub>-TPR results for NiO and Ni/ZrO<sub>2</sub> catalyst precursors.

Catalyst	H <sub>2</sub> Consumption ( $\mu\text{mol g}^{-1}$ )	e/Ni <sup>1</sup>	% NiO Species		
			$\alpha$	$\beta$	$\gamma$
NiO	1330	2.00	100	-	-
1.8Ni/Z <sub>m</sub>	31.04	2.02	-	60	40
5.1Ni/Z <sub>m</sub>	97.07	2.03	89	4	7
1.7Ni/Z <sub>hy</sub>	29.93	2.02	-	54	46
4.8Ni/Z <sub>hy</sub>	83.75	1.97	-	68	32

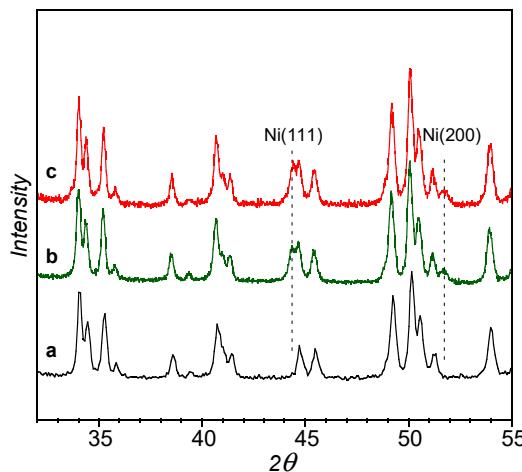
<sup>1</sup> Number of electrons consumed per Ni atom (e/Ni), calculated as (2-H<sub>2</sub> molecules consumed g<sup>-1</sup>)/(Ni atoms·g<sup>-1</sup>).

The characterization results for NiO/ZrO<sub>2</sub> catalyst precursors indicated that the use of amorphous Z<sub>hy</sub> as a starting material for the preparation of samples favoured the spreading of NiO on the support (FESEM and Raman evidence), particularly at high Ni contents. The higher number of sites available on Z<sub>hy</sub> than on Z<sub>m</sub> to anchor Ni<sup>2+</sup> species

during the impregnation step allowed the formation of smaller NiO particles on  $Z_{hy}$  than on  $Z_m$  during the subsequent thermal treatment, implying a higher number of Ni sites strongly interacting with  $Z_{hy}$  than with  $Z_m$  (TPR evidence).

### 3.1.3. Ni/ZrO<sub>2</sub> Catalysts

XRD patterns of Ni/ZrO<sub>2</sub> catalysts, irrespective of the different activation treatments, showed reflections assigned to the (111) and (200) planes of Ni metal (JCPDS card 04-0850) and to monoclinic ZrO<sub>2</sub> (JCPDS card 37-1484), as shown in Figure 5 for both 4.8Ni/ $Z_{hy}$  *ox-red* and *red*, as an example.



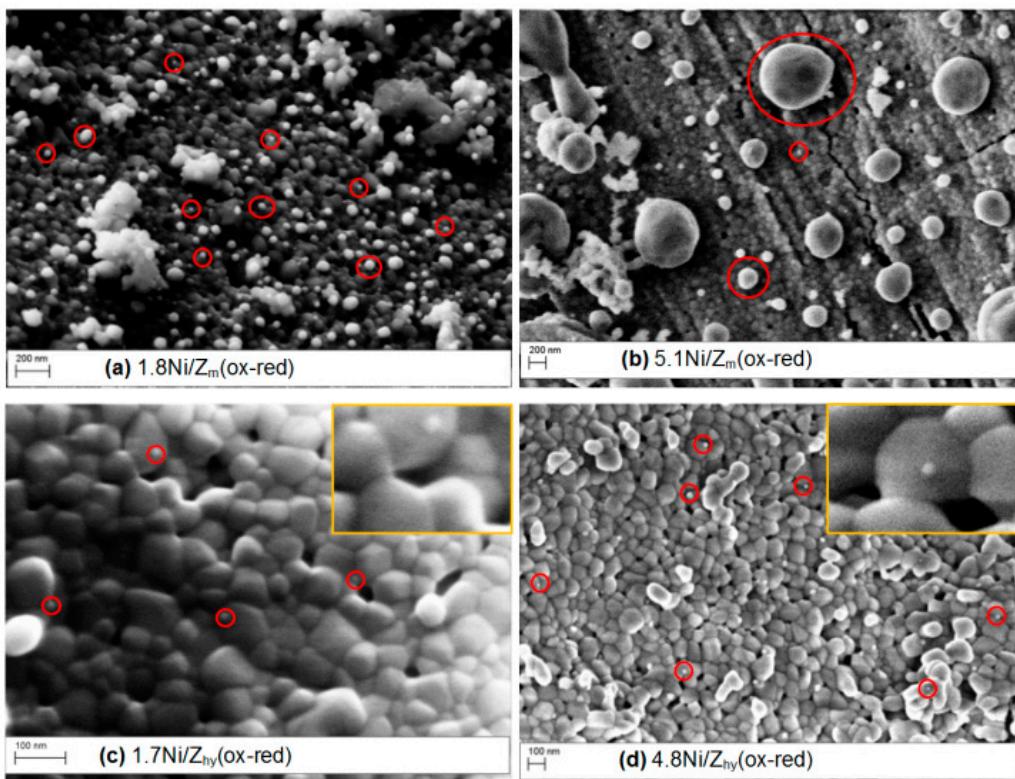
**Figure 5.** XRD patterns of monoclinic ZrO<sub>2</sub> and activated Ni/ZrO<sub>2</sub> catalysts. (a)  $Z_m$ ; (b) 4.8Ni/ $Z_{hy}$  (*red*); (c) 4.8Ni/ $Z_{hy}$  (*ox-red*).

Ni crystallite sizes, obtained by the Scherrer equation [41] using the (200) and/or (111) reflections, ranged from 18 to 35 nm for all of the samples (Table 1).

FESEM micrographs showed Ni particles larger in size with respect to XRD, indicating that, as for NiO, Ni crystallite aggregates were present on the zirconia surface. For Ni/ $Z_m$  catalysts, the Ni particle sizes (Table 2) were in the range of 5–75 nm in the 1.8Ni/ $Z_m$  (*ox-red*) sample (Figure 6a), and in the range of 30–1000 nm (highly heterogeneous in size) in the 5.1Ni/ $Z_m$  (*ox-red*) sample (Figure 6b). For Ni/ $Z_{hy}$  catalysts, small Ni particles of about 5–45 nm in diameter, in both the 1.7Ni/ $Z_{hy}$  (*ox-red*) and 4.8Ni/ $Z_{hy}$  (*ox-red*) samples, were detected (Figure 6c,d). Particle size values in the same ranges were found for the corresponding Ni/ $Z_{hy}$  (*red*) and Ni/ $Z_m$  (*red*) catalysts (Table 2).

Both the *ox-red* and *red* activation treatments caused a small particle sintering during the reductive process. In fact, instead of the expected decrease in particle size due to oxygen removal (based on NiO and Ni volume atomic densities,  $\rho_{NiO} = 2.69 \cdot 10^{22}$  Ni-atoms/cm<sup>3</sup> and  $\rho_{Ni} = 9.13 \cdot 10^{22}$  Ni-atoms/cm<sup>3</sup>, respectively) [45], we observed a small increase in size (compare  $d_{FESEM}$  of Ni with  $d_{FESEM}$  of NiO particles, Table 2). This result suggests that the NiO-ZrO<sub>2</sub> interaction in the catalyst precursors was strong enough to limit the aggregation of Ni particles during the reduction process. In agreement, metal particle sizes of the unsupported Ni metal sample ranged from 40 to 1500 nm, much larger than those detected on Ni/ZrO<sub>2</sub> samples (Table 2).

Overall, irrespective of the activation treatments, *ox-red* or *red*, Ni particles were smaller and more homogeneous in size in Ni/ $Z_{hy}$  than in Ni/ $Z_m$  catalysts, particularly at high Ni contents.



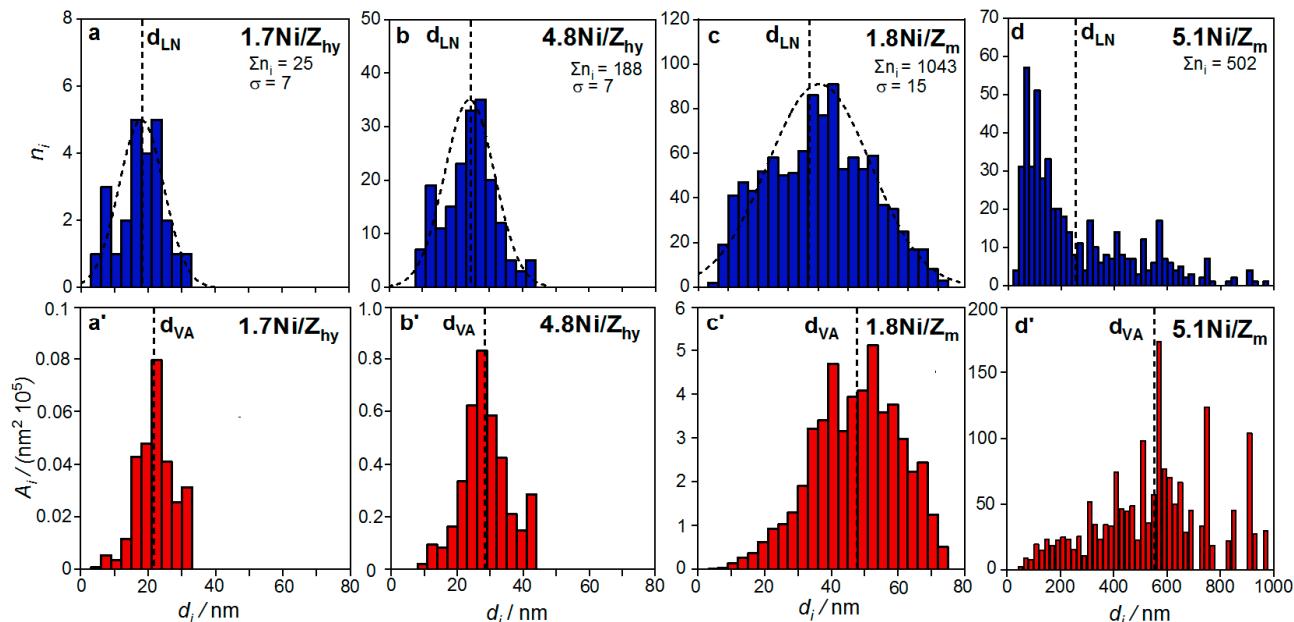
**Figure 6.** FESEM images of  $\text{Ni}/\text{Z}_{\text{hy}}$  (*ox-red*) and  $\text{Ni}/\text{Z}_m$  (*ox-red*) catalysts. Red circles and magnifications in sections (c) and (d) highlight some Ni particles.

### 3.1.4. Dispersion of Ni in the $\text{Ni}/\text{ZrO}_2$ Catalysts

To interpret the catalytic results (see below), the determination of Ni dispersion values is an important step in shedding light on the nature of active sites. The most frequently used method to evaluate the dispersion of the active phase is selective hydrogen chemisorption that allows the metal surface area measuring the  $\text{H}_2$  uptake to be determined. However, even in the samples with high Ni content and small particle size (as revealed by FESEM), the expected  $\text{H}_2$  uptake is too low to be confidently detected with our laboratory facilities (TPD and/or  $\text{H}_2$  chemisorption). As an alternative, the Ni dispersion could be evaluated by a physical technique able to measure particle diameters,  $d_i$ . Among the electron microscopy techniques, TEM is considered the most reliable and accurate physical technique for  $d_i$  measurements due to a resolution as high as 0.1 nm, whereas FESEM microscopy has a lower sensitivity limit (5 nm). As the particle sizes in our samples (Table 2) were above 5 nm in size, we confidently used FESEM microscopy to measure  $d_i$  values.

From FESEM image processing, we obtained a distribution of Ni-particle size,  $d_i$ , which was nearly symmetrical for the  $\text{Ni}/\text{Z}_{\text{hy}}$  samples and the  $1.8\text{Ni}/\text{Z}_m$  sample (Figure 7a–c), whereas it was asymmetrical and spread over a broad range of  $d_i$  sizes (up to 1000 nm) for the  $5.1\text{Ni}/\text{Z}_m$  sample (Figure 7d). From these distributions, the length-number mean diameter,  $d_{LN}$ , could be calculated (Table 2), as  $d_{LN} = \sum d_i / \sum n_i$  [45], where  $n_i$  is the number of metal particles with a specific diameter  $d_i$ . As reported in the literature [45,52], for phenomena depending on the particle surface, such as the catalytic activity, the mean diameter value that gives a better account of the total metal surface area is the volume/area mean diameter,  $d_{VA}$  ( $d_{VA} = \sum n_i d_i^3 / \sum n_i d_i^2 = 6 \cdot \sum n_i V_i / \sum n_i A_i$ ), which is directly related to the metal surface area [45]. In fact, in a particle-size distribution, large particles give the major contribution to the total metal surface area, as shown in the particle-surface distribution of  $\text{Ni}/\text{ZrO}_2$  samples at increasing particle size (Figures 7a'–d'). Therefore, the dispersion was calculated from the  $d_{VA}$  values, which are always higher

than the  $d_{LN}$  values [53]. The difference between  $d_{VA}$  and  $d_{LN}$  increased as the distribution broadened towards larger  $d_i$  sizes. In agreement, for the 1.7Ni/Z<sub>hy</sub>, 1.8Ni/Z<sub>m</sub> and 4.8Ni/Z<sub>hy</sub> samples,  $d_{VA}$  values were similar to the  $d_{LN}$  values, whereas for the 5.1Ni/Z<sub>m</sub> sample, the  $d_{VA}$  value was about twice that of the  $d_{LN}$  value (Figure 7 and Table 2). FESEM image processing gave similar Ni dispersion values for the 1.7Ni/Z<sub>hy</sub> and 4.8Ni/Z<sub>hy</sub> samples (about 4%), slightly lower for the 1.8Ni/Z<sub>m</sub> sample (about 2%), and about 20 times lower for the 5.1Ni/Z<sub>m</sub> sample (Table 2).

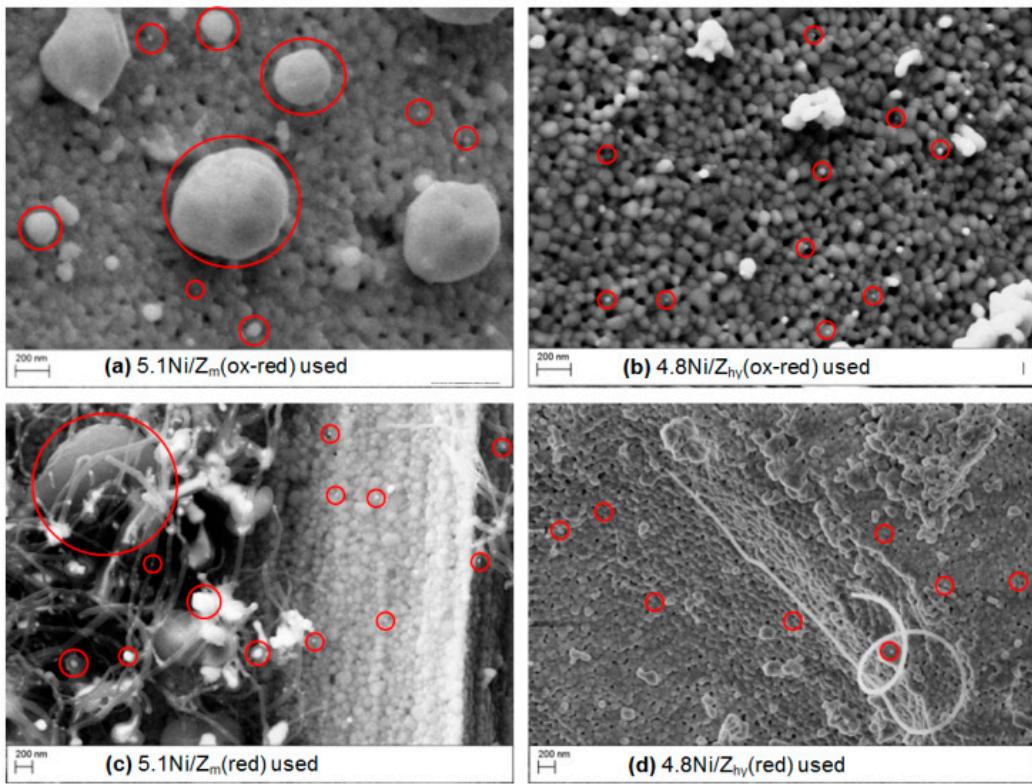


**Figure 7.** Ni-particle number ( $n_i$ ) (a–d) and Ni-particle surface ( $A_i = \pi n_i d_i^2$ ) (a'–d') distributions as a function of Ni-particle diameters ( $d_i$ ) for 1.7Ni/Z<sub>hy</sub> (a,a'), 4.8Ni/Z<sub>hy</sub> (b,b'), 1.8Ni/Z<sub>m</sub> (c,c') and 5.1Ni/Z<sub>m</sub> (d,d') samples. Dotted lines identify the length-number mean diameter ( $d_{LN}$ ) or volume-area mean diameter ( $d_{VA}$ ) values;  $\Sigma n_i$  stands for the total number of Ni particles from FESEM images. For samples showing a normal distribution curve of Ni particles size, the curve and the corresponding standard deviation  $\sigma$  are also reported.

### 3.1.5. Ni/ZrO<sub>2</sub> Used Catalysts

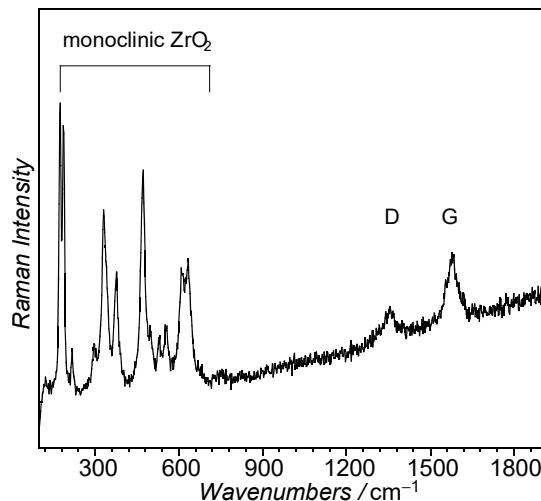
XRD patterns of all the used catalysts were similar to those of activated catalysts, showing the presence of both monoclinic ZrO<sub>2</sub> and Ni metal phases. The Ni mean crystallite size, estimated by XRD (Table 1), remained nearly unchanged under catalytic conditions.

FESEM analysis confirmed that Ni particles remained unchanged both in size and distribution after catalytic runs (Table 2), as revealed by comparing Figure 6b,d with Figure 8a,b for high-loaded catalysts as an example. These results supported the idea that, irrespective of dispersion, Ni metal particles did not undergo sintering under catalytic conditions. Carbon whiskers encapsulating Ni particles were detected on both 5.1Ni/Z<sub>m</sub> (red) and 4.8Ni/Z<sub>hy</sub> (red) used catalysts, whereas they were not observed on the Ni/Z (ox-red) used catalysts (Figure 8). This suggests that Ni particles were firmly anchored to the ZrO<sub>2</sub> support in Ni/Z (ox-red) catalysts, whereas in Ni/Z (red) catalysts, some Ni species showed a different reactivity that led to the growth of C whiskers incorporating Ni.



**Figure 8.** (a–d) FESEM images of 4.8Ni/Z<sub>hy</sub> and 5.1Ni/Z<sub>m</sub> used catalysts. Red circles highlight some Ni particles.

Raman spectroscopy, in agreement with FESEM results, clearly identified carbonaceous deposits on both 5.1Ni/Z<sub>m</sub> (*red*) and 4.8Ni/Z<sub>hy</sub> (*red*) used catalysts. As illustrated in Figure 9 for the 5.1Ni/Z<sub>m</sub> (*red*) used sample as an example, in addition to the bands due to monoclinic ZrO<sub>2</sub>, bands at about 1320 cm<sup>-1</sup> (D band) and 1584 cm<sup>-1</sup> (G band) were recorded. The G band ('graphite peak,' E<sub>2g</sub> symmetry) is characteristic of a perfect graphite lattice, whereas the D band ('defect peak,' A<sub>1g</sub> symmetry) is related to a defective graphitic lattice [53]. The nonoverlapping of the D and G bands is consistent with the presence of a structured carbon; the band at about 1584 cm<sup>-1</sup> could originate from graphitic carbon types on the catalyst's surface and that at about 1320 cm<sup>-1</sup> could originate from carbon nanoparticles or defective filamentous carbon. The relative intensity of the D band to the G band, I<sub>D</sub>/I<sub>G</sub>, has been taken as a measure of the structural disorder of the carbon material, a high ratio indicating a large amount of disorder in the structure and more reactive carbon species [24]. An I<sub>D</sub>/I<sub>G</sub> value of 0.9 was obtained by the curve fitting procedure for the 5.1Ni/Z<sub>m</sub> (*red*) used sample, indicating the presence of a significant contribution of graphite-type structures.



**Figure 9.** Raman spectrum of the 5.1Ni/Z<sub>m</sub> (*red*) used catalyst. D and G identify disordered-graphite and graphite bands, respectively.

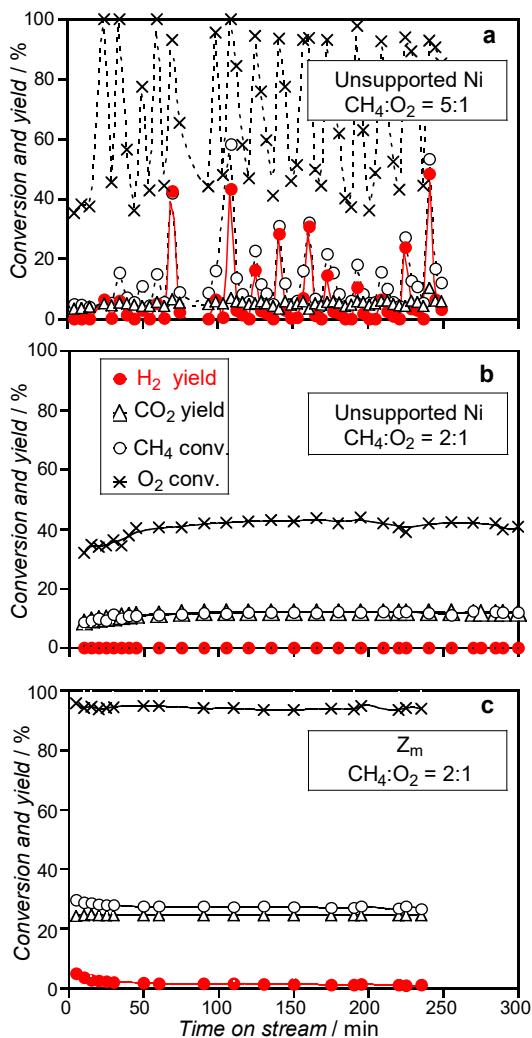
Overall, characterization results of catalysts, before and after catalytic runs, showed that the starting material adopted for the preparation, Z<sub>hy</sub> or Z<sub>m</sub>, (i) affected the Ni particle sizes (smaller for the high-surface Z<sub>hy</sub>) because it affected the NiO particle sizes in the catalyst precursors; (ii) it did not affect the strength of the interaction between Ni particles and the ZrO<sub>2</sub> surface. In all the catalysts, the activation treatment of the NiO/ZrO<sub>2</sub> precursors did not affect the Ni particle sizes or the Ni-ZrO<sub>2</sub> interactions that were strong enough to prevent the sintering of the Ni metal particles.

### 3.2. Catalytic Activity

Unsupported Ni as foil or wires was reported to be somewhat active for the CH<sub>4</sub>-CPO using a methane-rich mixture CH<sub>4</sub>:O<sub>2</sub> = 5:1, with a peculiar oscillatory activity [54,55]. Unsupported Ni metal powder was tested for the CH<sub>4</sub>-CPO using the methane-rich mixture and the stoichiometric mixture (CH<sub>4</sub>:O<sub>2</sub> = 2:1). Irrespective of the activation procedure (*ox-red* or *red*), Ni metal powder showed an oscillatory on-off activity at 1023 K with the methane-rich mixture (Figure 10a), whereas it was inactive for syngas production with the stoichiometric mixture, showing only activity for the CH<sub>4</sub> total combustion (Figure 10b).

The Z<sub>m</sub> support was inactive for the CH<sub>4</sub>-CPO in the temperature range of 773–1023 K with the CH<sub>4</sub>:O<sub>2</sub> = 2:1 mixture and, irrespective of the activation treatment, yielded a total oxidation of CH<sub>4</sub> above 823 K (maximum CO<sub>2</sub> yield of about 25% at 1023 K, Figure 10c).

Unlike ZrO<sub>2</sub> and unsupported Ni metal, Ni/ZrO<sub>2</sub> catalysts were active for the CH<sub>4</sub>-CPO, indicating that the metal–support interaction plays a key role in developing Ni active species. However, the catalytic behaviour markedly depended on the activation procedure (*ox-red* or *red*), as illustrated in the following sections.

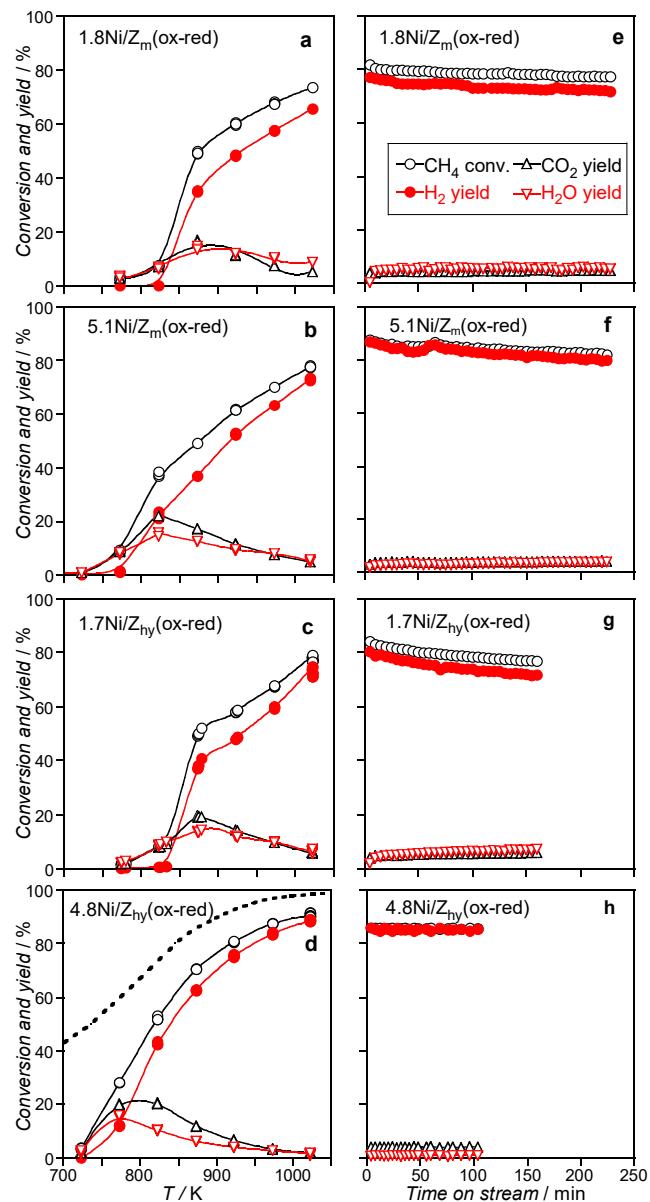


**Figure 10.** Catalytic activity of unsupported Ni metal and monoclinic  $\text{ZrO}_2$  ( $Z_m$ ) with reactant mixtures at different  $\text{CH}_4:\text{O}_2$  ratios.  $\text{CH}_4$  and  $\text{O}_2$  conversions and  $\text{H}_2$  and  $\text{CO}_2$  yields as a function of time on stream at 1023 K for *ox-red*-activated catalysts. (a):  $[\text{CH}_4] = 2\%$ ,  $[\text{O}_2] = 0.4\%$ . (b,c):  $[\text{CH}_4] = 2\%$ ,  $[\text{O}_2] = 1\%$ ,  $\text{N}_2$  as balance; total flow rate =  $150 \text{ cm}^3 (\text{STP}) \cdot \text{min}^{-1}$ .

### 3.2.1. Catalysts Activated by Oxidation–Reduction Treatment

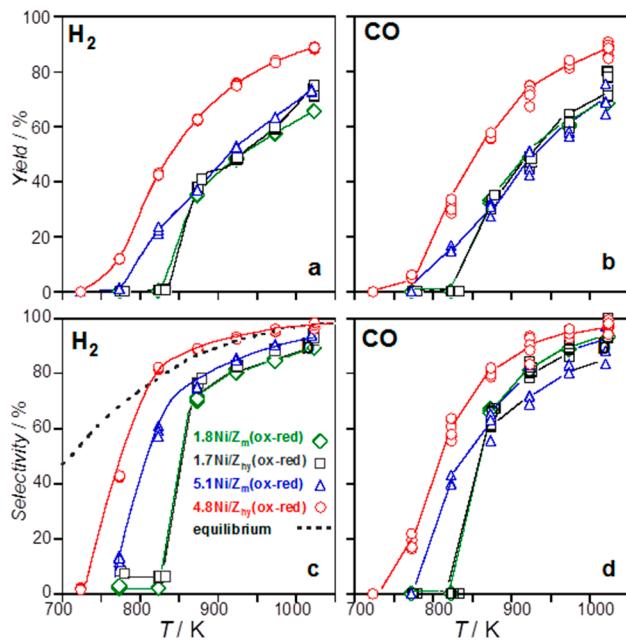
All the  $\text{Ni}/\text{Z}_m$  (*ox-red*) and  $\text{Ni}/\text{Z}_{hy}$  (*ox-red*) catalysts were highly active (Figures 11a–d and 12a,b) and selective (Figure 12c,d) for the  $\text{CH}_4$ -CPO at 1023 K with the  $\text{CH}_4:\text{O}_2 = 2:1$  mixture. The most active 4.8Ni/ $\text{Z}_{hy}$  (*ox-red*) sample reached, at 1023 K, a  $\text{CH}_4$  conversion and  $\text{H}_2$  selectivity close to the equilibrium (90% and 95%, respectively [14], Figures 11d and 12c).

For all the catalysts, decreasing the temperature below 1023 K decreased the  $\text{H}_2$  selectivity ( $\text{H}_2$  yield lower than  $\text{CH}_4$  conversion), indicating the occurrence of side reactions. The formation of similar amounts of  $\text{CO}_2$  and  $\text{H}_2\text{O}$  as by-products suggested the occurrence of  $\text{CH}_4$  total combustion, with a volcano-shape activity trend as a function of temperature (maximum  $\text{CO}_2$  yield of 20%). Below 923 K, more evident for the concentrated 5.1Ni/ $\text{Z}_m$  (*ox-red*) and 4.8Ni/ $\text{Z}_{hy}$  (*ox-red*) samples, the  $\text{H}_2\text{O}$  yield was slightly lower than the  $\text{CO}_2$  yield, and the  $\text{H}_2$  yield was slightly higher than the  $\text{CO}$  yield, suggesting that the WGS reaction occurred to a small extent. For all the catalysts, the activity and selectivity for the  $\text{CH}_4$ -CPO dropped to zero, further lowering the temperature, and they were completely restored by a subsequent temperature increase, indicating that side-reactions causing irreversible deactivation did not occur.



**Figure 11.** Catalytic activity ( $\text{CH}_4$  conversion and  $\text{H}_2$ ,  $\text{CO}_2$  and  $\text{H}_2\text{O}$  yields) of  $\text{Ni}/\text{ZrO}_2$  catalysts after *ox-red* activation treatment in experiments as a function of temperature (left side: (a–d)) or as a function of time on stream at 1023 K (right side: (e–h)). Reactant mixture:  $[\text{CH}_4] = 2\%$ ,  $[\text{O}_2] = 1\%$ ,  $\text{N}_2$  as balance; total flow rate =  $150 \text{ cm}^3 (\text{STP}) \cdot \text{min}^{-1}$ . Dotted line:  $\text{CH}_4$  conversion for the same mixture calculated at thermodynamic equilibrium (data from Ref. [11]).

The activity and selectivity as a function of time on stream at a set temperature were stable for the  $4.8\text{Ni}/\text{Z}_{\text{hy}}$  (*ox-red*) catalyst, but were slightly decreased for all the other catalysts (Figure 11e–h). The reproducibility of activity and selectivity in the whole temperature range was observed in each run and in consecutive runs (Figure S3a), with the error affecting the conversion values of, at maximum, 7%. All of these results suggest that no significant changes in the amount and/or type of active Ni sites occurred during the reaction and that the slight decrease in activity with time possibly arose from the deposition of a very small amount of coke on Ni active sites, in agreement with the literature [23].



**Figure 12.** Comparison of activity and selectivity for Ni/ZrO<sub>2</sub> catalysts after *ox-red* activation treatment. H<sub>2</sub> yield and selectivity (**a,c**) and CO yield and selectivity (**b,d**) as a function of temperature. Reactant mixture: [CH<sub>4</sub>] = 2%, [O<sub>2</sub>] = 1%, N<sub>2</sub> as balance; total flow rate = 150 cm<sup>3</sup> (STP)·min<sup>-1</sup>. Dotted line: H<sub>2</sub> selectivity for the same mixture calculated at thermodynamic equilibrium (data from Ref. [11]).

Similar H<sub>2</sub> and CO yields (Figure 12a,b) and selectivities (Figure 12c,d) were observed for dilute 1.8Ni/Z<sub>m</sub> (*ox-red*) and 1.7Ni/Z<sub>hy</sub> (*ox-red*) catalysts, indicating that at low Ni loading, the starting supports allowed the formation of similar Ni active sites. The increase in Ni loading caused a different behaviour of catalysts prepared using Z<sub>m</sub> or Z<sub>hy</sub> starting materials. Specifically, for Ni/Z<sub>m</sub> (*ox-red*) catalysts, the increase in Ni loading lowered the reaction light-off of about 50 K, leaving the H<sub>2</sub> and CO yield and selectivity values unchanged above the light-off temperature. By contrast, for Ni/Z<sub>hy</sub> (*ox-red*) catalysts, the increase in Ni loading caused a more consistent lowering of the reaction light-off temperature (about 100 K), giving higher H<sub>2</sub> and CO yield and selectivity values for 4.8Ni/Z<sub>hy</sub> (*ox-red*) than for 1.7Ni/Z<sub>hy</sub> (*ox-red*) catalysts (Figure 12a–d). Therefore, the 4.8Ni/Z<sub>hy</sub> (*ox-red*) catalyst showed the best performances, exhibiting the lowest CPO light-off temperature and the highest H<sub>2</sub> and CO yield and selectivity values.

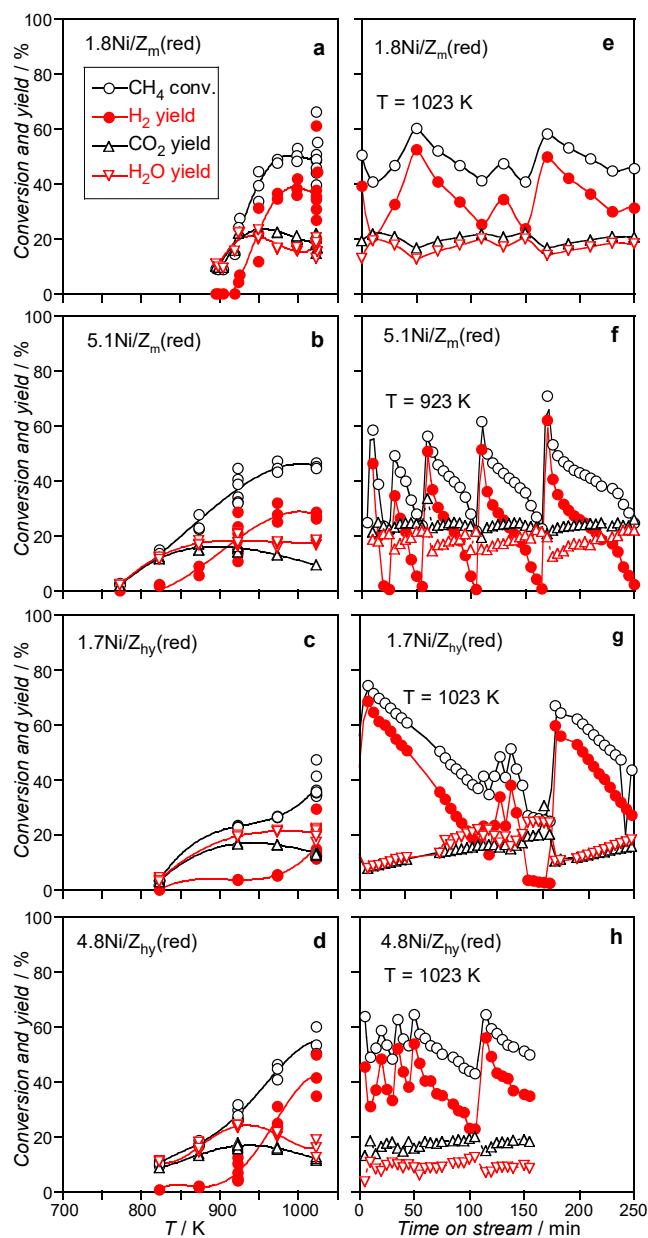
As 1.8Ni/Z<sub>m</sub> (*ox-red*) and 1.7Ni/Z<sub>hy</sub> (*ox-red*) catalysts showed similar activity, the starting material Z<sub>hy</sub> or Z<sub>m</sub> used for the impregnation seemed to have no effect on the active site reactivity. The evidence that the 4.8Ni/Z<sub>hy</sub> (*ox-red*) catalyst with the smallest Ni particles had an activity higher than that of the 5.1Ni/Z<sub>m</sub> (*ox-red*) catalyst with the largest particles (FESEM evidence, Table 2) might raise the expectation that Ni dispersion favours catalytic activity.

To verify whether dispersion can be the key parameter affecting activity, the dependence of the rate of H<sub>2</sub> production (R<sub>H2</sub>/molecules s<sup>-1</sup>·g<sup>-1</sup>) on the number of exposed Ni atoms, N<sub>Ni(exp)</sub> (Ni(exp) atoms·g<sup>-1</sup>), was analysed for catalysts with different Ni dispersions (Table 2). The N<sub>Ni(exp)</sub> value of the 4.8Ni/Z<sub>hy</sub> (*ox-red*) catalyst (1.8·10<sup>19</sup> atoms·g<sup>-1</sup>) was about 20 times higher than that of the 5.1Ni/Z<sub>m</sub> (*ox-red*) catalyst (9.4·10<sup>17</sup> atoms·g<sup>-1</sup>), whereas the rate R<sub>H2</sub> at 823 K for the 4.8Ni/Z<sub>hy</sub> (*ox-red*) catalyst was only about twice as high than that for the 5.1Ni/Z<sub>m</sub> (*ox-red*) catalyst (2.2·10<sup>19</sup> vs. 1.2·10<sup>19</sup> molecules·s<sup>-1</sup>·g<sup>-1</sup>, respectively). The lack of a linear correlation between R<sub>H2</sub> and N<sub>Ni(exp)</sub> indicated that neither all nor a constant fraction of exposed Ni atoms were the active sites, only Ni species with peculiar properties. As a consequence, neither the different Z<sub>hy</sub> or Z<sub>m</sub> starting material

alone nor Ni dispersion alone was the factor affecting the activity of Ni sites for CH<sub>4</sub>-CPO.

### 3.2.2. Catalysts Activated by Reduction Treatment

All the Ni/Z<sub>m</sub> (*red*) and Ni/Z<sub>hy</sub> (*red*) catalysts were less active (Figure 13a–d) and selective (Figure S4a,b) for the CH<sub>4</sub>-CPO than the corresponding *ox-red* samples in the whole temperature range. They gave scattered values of H<sub>2</sub> and CO yields and no scattered combustion by-product values. Both the scattered H<sub>2</sub> and CO yields and the lower activity of *red* samples with respect to the *ox-red* samples suggested that the *red* activation treatment generated active sites in a lower amount and with features unlike those obtained by the *ox-red* procedure.



**Figure 13.** Catalytic activity (CH<sub>4</sub> conversion and H<sub>2</sub>, CO<sub>2</sub> and H<sub>2</sub>O yields) of Ni/ZrO<sub>2</sub> catalysts after *red* activation treatment in experiments as a function of temperature (left side: a–d) or as a function of time on stream at set temperatures (right side: e–h). Reactant mixture: [CH<sub>4</sub>] = 2%, [O<sub>2</sub>] = 1%, N<sub>2</sub> as balance; total flow rate = 150 cm<sup>3</sup> (STP)·min<sup>-1</sup>.

For all Ni/Z<sub>m</sub> (*red*) and Ni/Z<sub>hy</sub> (*red*) catalysts, the scattered trend of activity was reproduced in subsequent runs (Figure S3b), indicating that, as for Ni/Z<sub>m</sub> (*ox-red*) and Ni/Z<sub>hy</sub> (*ox-red*) catalysts, no irreversible modification of the active Ni sites occurred during catalytic runs.

By investigating the activity at a set temperature, the Ni/Z<sub>m</sub> (*red*) and Ni/Z<sub>hy</sub> (*red*) catalysts exhibited CH<sub>4</sub> conversion and H<sub>2</sub> and CO yields markedly changing as a function of time on stream with *in-phase saw-tooth shape* oscillations (Figure 13e–h), indicating an oscillating behaviour of the CH<sub>4</sub>-CPO reaction. In the oscillating cycles, the total oxidation side-reaction occurred for all the catalysts, as evidenced by the CH<sub>4</sub> conversion always being higher than the H<sub>2</sub> yield and H<sub>2</sub>O and CO<sub>2</sub> formation.

WGS side reactions also occurred to a small extent, as the H<sub>2</sub> yield was slightly higher than that of CO, and the CO<sub>2</sub> yield was slightly higher than that of H<sub>2</sub>O. The slightly oscillating behaviour of CO<sub>2</sub> and H<sub>2</sub>O yields (Figure 13e–h), opposite in phase to that of the CH<sub>4</sub>-CPO, suggested an oscillating behaviour even for the WGS reaction. The oscillation of the WGS reaction might arise from the significant change in CO concentration as the reactant, consequent to the CH<sub>4</sub>-CPO oscillation. As the change in the relative amount of CO<sub>2</sub> and H<sub>2</sub>O yields was small and arose from WGS oscillation, the amounts of CO<sub>2</sub> and H<sub>2</sub>O arising from CH<sub>4</sub> combustion were constant, indicating that the CH<sub>4</sub> total oxidation did not oscillate. In addition, as the CO<sub>2</sub> yield in the Ni/Z (*red*) catalysts was roughly similar to that observed for the Z<sub>m</sub> support (see Figure 10c), we suggest that the active sites yielding CH<sub>4</sub> total oxidation are support sites, and those yielding oscillating CH<sub>4</sub>-CPO and WGS behaviours are Ni sites.

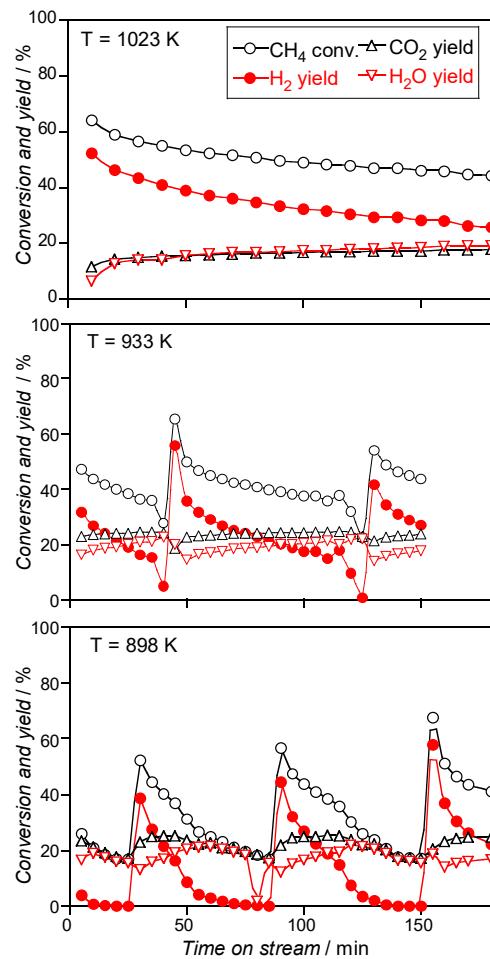
The remarkable oscillation of the CH<sub>4</sub>-CPO reaction for the Ni/Z<sub>m</sub> (*red*) and Ni/Z<sub>hy</sub> (*red*) catalysts can be tentatively explained according to evidence widely discussed in the literature for various metal or supported metal systems [55–62]. For supported noble metal catalysts (Ru and Pd), the oscillating behaviour was ascribed to a periodical oxidation–reduction of the active metal sites, with the metal species being active for the CH<sub>4</sub>-CPO and the oxidised species being active for the CH<sub>4</sub> combustion [58,61]. In particular, an oxidation/reduction front of Pd nanoparticles was suggested to move through the catalytic fixed-bed, yielding the periodical predominance of the CH<sub>4</sub> total oxidation with respect to the partial oxidation [61]. An *operando* XRD study on a Ni foil showed unambiguously that the oscillations of the CH<sub>4</sub>-CPO reaction were due to the reversible oxidation of Ni metal to NiO [63]. For Ni and Co foils [60] and Ni wires [56], it was suggested that the reversible oxidation–reduction of surface sites depended on the gas-phase composition and on changes in the local temperature due to the exothermicity of the reactions at the metal surface.

In agreement with all these findings, we suggest that in both Ni/Z<sub>m</sub> (*red*) and Ni/Z<sub>hy</sub> (*red*) catalysts, the CH<sub>4</sub>-CPO oscillating behaviour arose from the periodical oxidation–reduction of some Ni sites, which were CPO-inactive when oxidised and CPO-active when reduced. Unlike supported Ru and Pd catalysts [58,61], the oxidised Ni species that formed during the oscillation cycle contributed little to the CH<sub>4</sub> total oxidation, as in our catalysts, CH<sub>4</sub> total oxidation did not oscillate.

The oscillating behaviour depended on temperature, contact time and O<sub>2</sub> content in the mixture. Specifically, (i) the trends of conversions and yields changed form or disappeared, depending on temperature (Figure 14), as the surface reactions causing changes in the Ni oxidation state depended on temperature by different activation energy values; (ii) oscillations disappeared, increasing the total flow rate (Figure S5), in agreement with the literature, showing that short contact times favoured CH<sub>4</sub>-CPO compared to total oxidation [63]; (iii) oscillations appeared when the O<sub>2</sub> content was above 0.5% (Figure S6), as a higher oxygen coverage of the metal surface favoured the oxidation of Ni sites [60].

To explain why Ni/Z<sub>m</sub> (*red*) and Ni/Z<sub>hy</sub> (*red*) catalysts yielded oscillations, whereas Ni/Z<sub>m</sub> (*ox-red*) and Ni/Z<sub>hy</sub> (*ox-red*) did not, we rationalized our results considering carbon balance, oscillation reproducibility, and Ni particle size for the used samples. As, during

the oscillation cycles, the carbon balance was satisfactory, the oscillating behaviour could not be related to the process of the covering/restoring of active Ni sites (carbon deposition followed by auto-thermally induced carbon oxidation). In addition, as the oscillating behaviour was the same in subsequent runs, the number and features of Ni sites causing oscillations were not affected by aging in the CH<sub>4</sub>-CPO feed and by the re-activation treatment between two consecutive runs. Recalling that for all the catalysts, the *red* activation treatment yielded Ni metal particles similar in size to those obtained by the *ox-red* activation treatment (see Table 2), the oscillating activity of catalysts could not be attributed to Ni particles of specific size. Therefore, we suggest that the *red* activation treatment yielded specific Ni sites able to easily change their oxidation state, inducing an oscillating process.



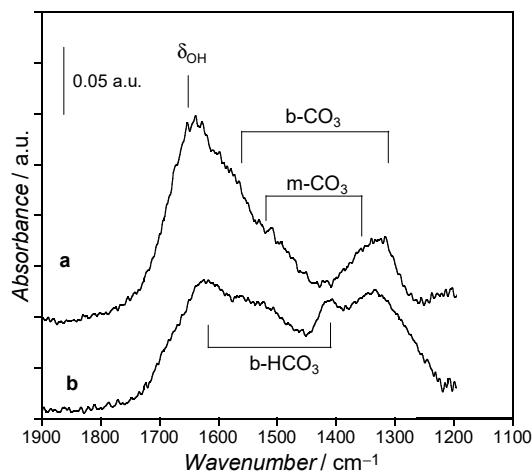
**Figure 14.** CH<sub>4</sub> conversion and H<sub>2</sub>, CO<sub>2</sub> and H<sub>2</sub>O yields on 5.1Ni/Z<sub>m</sub> (*red*) catalyst as a function of time on stream for specific temperatures. Reactant mixture: [CH<sub>4</sub>] = 2%, [O<sub>2</sub>] = 1%, N<sub>2</sub> as balance; total flow rate = 150 cm<sup>3</sup> (STP)·min<sup>-1</sup>.

### 3.3. Surface Species Evolution during Activation Treatments by *in situ* FTIR

To gain an insight into the Ni surface species that might be involved in the oscillating activity of Ni/Z<sub>m</sub> (*red*) and Ni/Z<sub>hy</sub> (*red*) catalysts, *in situ* FTIR analysis was performed in a gas flow simulating the activation treatments of the catalyst precursors.

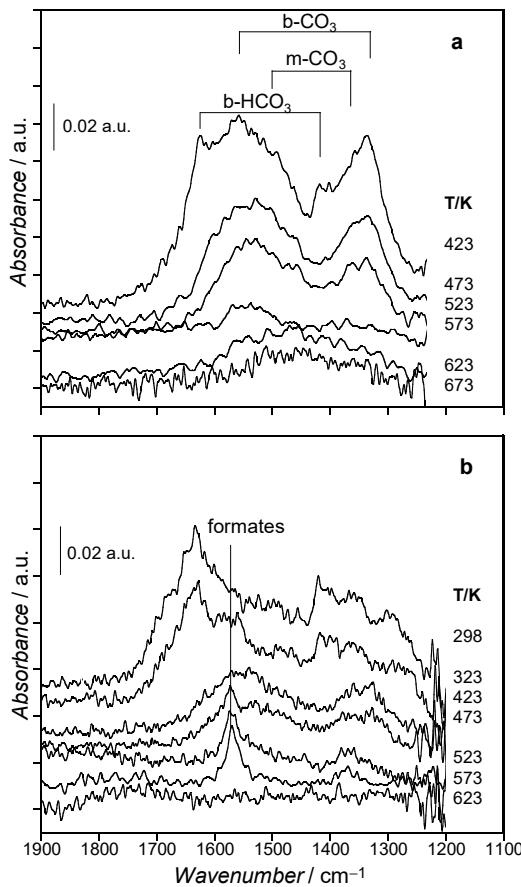
The spectra of the Z<sub>m</sub> support and the 4.8NiO/Z<sub>hy</sub> precursor recorded in air at 298 K (Figure 15) showed a broad band at about 3600 cm<sup>-1</sup> due to H-bonded OH (ν<sub>OH</sub>, spectral region not shown) and an envelope in the 1650–1250 cm<sup>-1</sup> region, containing bands characteristic of adsorbed water (at about 1630 cm<sup>-1</sup>, δ<sub>OH</sub> mode), monodentate carbonates (m-CO<sub>3</sub>, at 1475 and 1355 cm<sup>-1</sup>), bidentate carbonates (b-CO<sub>3</sub>, at 1560 and 1330 cm<sup>-1</sup>) and

bidentate bicarbonates ( $\text{b-HCO}_3$ , at 1690 and  $1410\text{ cm}^{-1}$ ) adsorbed on the zirconia surface [64–68].



**Figure 15.** FTIR spectra of surface species recorded in air at 298 K. (a)  $\text{Z}_\text{m}$  support; (b) 4.8NiO/ $\text{Z}_\text{hy}$  catalyst precursor.

The exposure of the 4.8NiO/ $\text{Z}_\text{hy}$  precursor to an  $\text{O}_2/\text{N}_2$  flow at increasing temperature, simulating the oxidative step of the *ox-red* activation procedure, produced a progressive decrease in intensity of the bands due to adsorbed  $\text{H}_2\text{O}$ , carbonates and bicarbonates, until their disappearance above 573 K (Figure 16a).



**Figure 16.** In situ FTIR spectra of surface species during activation treatments on 4.8NiO/ $\text{Z}_\text{hy}$  catalyst precursor. Spectra recorded in  $\text{O}_2/\text{N}_2$  oxidative flow (a), and in  $\text{H}_2/\text{N}_2$  reductive flow (b) at increasing temperature.

The exposure of the 4.8NiO/Z<sub>hy</sub> precursor to an H<sub>2</sub>/N<sub>2</sub> flow, at increasing temperature, simulating the *red* activation procedure, produced a progressive decrease in intensity of adsorbed H<sub>2</sub>O and carbonates/bicarbonates bands. At 473 K, a new band at about 1575 cm<sup>-1</sup> appeared, whose intensity reached a maximum at 573 K and then decreased with temperature. All the bands disappeared above 623 K (Figure 16b). The band at 1575 cm<sup>-1</sup> was assigned to formate species [69], arising from the reduction of the adsorbed carbonates/bicarbonates surface species in the H<sub>2</sub> flow.

The formation of reduced carbon-containing species solely after the *red* activation treatment of the precursors could account for the dependence of the catalytic behaviour on the activation procedures. During the *red* activation, carbonates species adsorbed on the support could be reduced to formates that, at higher temperature, are supposed to form carbide-like species (C<sup>δ-</sup>), not detectable by IR. The C<sup>δ-</sup> species located on the ZrO<sub>2</sub> surface in proximity of the Ni particles could induce a positive charge on Ni atoms at the boundary between metal particles and the support, yielding Ni<sup>δ+</sup> surface species. These species could easily change their oxidation state, thus favouring the oscillating behaviour of the CH<sub>4</sub>-CPO reaction. As the oscillating catalytic behaviour was retained by subsequent *red-ox* cycles, it can be assumed that the C<sup>δ-</sup> species, once formed, are stable on the surface of *red*-activated catalysts. The formation of carbide-like species in Ni/Z<sub>m</sub> (*red*) and Ni/Z<sub>hy</sub> (*red*) catalysts, hypothesized based on FTIR results, can find some correspondence in the presence of carbon nanotubes observed by FESEM. In fact, by consensus, the presence of nickel carbide species was well recognised to be involved in the induction step of the carbon nanotubes formation on Ni metal particles [70]. Conversely, when the catalysts precursors were activated by the *ox-red* procedure, adsorbed carbonates desorbed during the oxidation step, and carbide-like species could not form during the reduction step. Consequently, Ni active sites with a pure metal character formed, thus preventing the oscillating phenomenon.

#### 3.4. Hypothesis on Active Sites for the CH<sub>4</sub>-CPO

The activity of both Ni/Z<sub>m</sub> (*ox-red*) and Ni/Z<sub>hy</sub> (*ox-red*) catalysts, much higher than that of unsupported Ni, and the oscillating behaviour of Ni/Z<sub>m</sub> (*red*) and Ni/Z<sub>hy</sub> (*red*) catalysts suggest that the active sites are possibly Ni sites at the boundary between metal particles and the support. The electron density of these Ni sites is sufficiently affected by the interaction with zirconia and/or with nearby carbide-like species to modify the catalytic behaviour of these Ni sites with respect to unsupported Ni metal particles.

Furthermore, Ni active sites should only be a very small fraction of the boundary species, as TPR results indicated a markedly lower amount of interacting Ni species in the 5.1Ni/Z<sub>m</sub> catalyst than in the 4.8Ni/Z<sub>hy</sub> catalyst (11% vs 100%), both showing similar catalytic performance.

In the literature, the CH<sub>4</sub>-CPO mechanism was described as a direct oxidation route, requiring several Ni sites able to activate methane and oxygen in adjacent positions (Ni–C, Ni–O and Ni–H) [14,17,36,60]. According to this pathway, we suggest that the active sites for the CH<sub>4</sub>-CPO on Ni/ZrO<sub>2</sub> catalysts should be poly-nuclear metal sites at the metal particle boundary, in a specific configuration. A similar proposal on the presence of cooperating nearby sites has been made to explain the catalytic activity of Ni supported on grafted ZrO<sub>2</sub>-Al<sub>2</sub>O<sub>3</sub> catalysts [36]. The amount of poly-nuclear metal active sites was lower in the Ni/Z (*red*) than in the Ni/Z (*ox-red*) catalysts because, in the Ni/Z (*red*), some Ni<sup>δ+</sup> carbide-like species formed, inducing CPO oscillation.

## 4. Conclusions

Ni metal particles supported on monoclinic ZrO<sub>2</sub>, prepared via the impregnation of two different starting materials, were active for the partial oxidation of CH<sub>4</sub>. As both monoclinic ZrO<sub>2</sub> and unsupported Ni metal were inactive for the CH<sub>4</sub>-CPO under the same experimental conditions, the metal–support interaction played a key role in developing active Ni sites. Although the starting material strongly influenced the size of the

supported NiO particles due to the different number of available anchoring sites, the NiO-ZrO<sub>2</sub> interaction was strong enough to prevent the sintering of metal particles formed during the *ox-red* or *red* activation treatment. In the same way, the Ni-ZrO<sub>2</sub> interaction was strong enough to guarantee the stability of metal particles during the catalytic runs.

Different activation treatments of catalysts yielded different catalytic properties: the *ox-red* treatment yielded nonoscillating activity and a selectivity of Ni sites, whereas the *red* treatment induced an oscillating behaviour. It can be suggested that the *red* activation treatment reduced NiO and carbonates adsorbed on the zirconia surface, leading to the formation of Ni<sup>δ+</sup> carbide-like species. These Ni<sup>δ+</sup> sites, having a greater tendency to change the oxidation state, were responsible for oscillations in the catalytic activity. Conversely, the *ox-red* treatment, removing adsorbed carbonates during the oxidation step, yielded Ni active sites with metal characteristics, producing a nonoscillating catalytic activity.

We conclude that Ni dispersion was not the main factor affecting the activity and that the active sites were a small fraction of all the exposed Ni atoms, likely those at the boundary of the metal particles in a specific configuration and nuclearity, strongly depending on the activation procedure. The interaction of these specific Ni atoms with the zirconia surface and/or with nearby carbide-like species was strong enough to modify the catalytic behaviour of the Ni species with respect to the unsupported Ni metal.

**Supplementary Materials:** The following are available online at [www.mdpi.com/1996-1944/14/10/2495/s1](http://www.mdpi.com/1996-1944/14/10/2495/s1), Figure S1: Catalytic activity with contact time. Figure S2: Nitrogen adsorption/desorption isotherms (a) and pore size distributions (b) for the Z<sub>m</sub> support and NiO/ZrO<sub>2</sub> catalyst precursors. Figure S3: Reproducibility of activity in four subsequent runs on two representative samples activated by the *ox-red* treatment (a) or by the *red* treatment (b). Figure S4: Percent H<sub>2</sub> selectivity (a) and percent CO selectivity (b) for Ni/ZrO<sub>2</sub> catalysts after *red* activation treatment as a function of temperature. Figure S5: Catalytic activity of 5.1Ni/Z<sub>m</sub> (*red*) catalyst at 923 K with different total flow rates. CH<sub>4</sub> conversion and H<sub>2</sub>, CO<sub>2</sub> and H<sub>2</sub>O yields as a function of time on stream. Figure S6: Catalytic activity of 1.8Ni/Z<sub>m</sub> (*red*) catalyst at 1023 K with different O<sub>2</sub> contents in the feed. CH<sub>4</sub> conversion and H<sub>2</sub>, CO<sub>2</sub> and H<sub>2</sub>O yields as a function of time on stream.

**Author Contributions:** Conceptualization, D.P. and D.G.; data curation, D.P. and M.C.C.; formal analysis, S.T., M.C.C. and G.L.; funding acquisition, D.P. and D.G.; investigation, I.P., S.T., G.L., L.A. and I.L.; methodology, D.P.; project administration, D.P. and D.G.; resources, G.L. and L.A.; supervision, D.P. and D.G.; writing—original draft, G.L.; writing—review and editing, D.G., M.C.C., I.P. and D.P. All authors have read and agreed to the published version of the manuscript.

**Funding:** This research was funded by Sapienza University of Roma, grant number RM11715C7E58658B.

**Institutional Review Board Statement:** Not applicable.

**Informed Consent Statement:** Not applicable.

**Data Availability Statement:** Data are contained within the article and supplementary material.

**Acknowledgments:** The authors are grateful to Francesco Mura (Sapienza Nanotechnology and Nanoscience Laboratory (SNN-Lab), Sapienza University of Rome) for FESEM analyses.

**Conflicts of Interest:** The authors declare no conflict of interest. The funders had no role in the design of the study; in the collection, analyses, or interpretation of data; in the writing of the manuscript, or in the decision to publish the results.

## References

1. Faramawy, S.; Zaki, T.; Sakr, A.-E. Natural gas origin, composition, and processing: A review. *J. Nat. Gas. Sci. Eng.* **2016**, *34*, 34–54, doi:10.1016/j.jngse.2016.06.030.
2. Olajire, A.A. Valorization of greenhouse carbon dioxide emissions into value-added products by catalytic processes. *J. CO<sub>2</sub> Util.* **2013**, *3–4*, 74–92, doi:10.1016/j.jcou.2013.10.004.
3. Horn, R.; Schlägl, R. Methane Activation by Heterogeneous Catalysis. *Catal. Lett.* **2015**, *145*, 23–39, doi:10.1007/s10562-014-1417-z.

4. Pen˜aM.; Gómez, J.; Fierro, J. New catalytic routes for syngas and hydrogen production. *Appl. Catal. A: Gen.* **1996**, *144*, 7–57, doi:10.1016/0926-860x(96)00108-1.
5. Al-Sayari, S.A. Recent Developments in the Partial Oxidation of Methane to Syngas. *Open Catal. J.* **2013**, *6*, 17–28, doi:10.2174/1876214x20130729001.
6. Ghoneim, S.A.; El-Salamony, R.A.; El-Temtamy, S.A. Review on Innovative Catalytic Reforming of Natural Gas to Syngas. *World J. Eng. Technol.* **2016**, *4*, 116–139, doi:10.4236/wjet.2016.41011.
7. Aasberg-Petersen, K.; Dybkjær, I.; Ovesen, C.; Schjødt, N.; Sehested, J.; Thomsen, S. Natural gas to synthesis gas – Catalysts and catalytic processes. *J. Nat. Gas. Sci. Eng.* **2011**, *3*, 423–459, doi:10.1016/j.jngse.2011.03.004.
8. Rostrup-Nielsen, J. Steam reforming and chemical recuperation. *Catal. Today* **2009**, *145*, 72–75, doi:10.1016/j.cattod.2008.06.019.
9. Meloni, E.; Martino, M.; Palma, V. A Short Review on Ni Based Catalysts and Related Engineering Issues for Methane Steam Reforming. *Catalysts* **2020**, *10*, 352, doi:10.3390/catal10030352.
10. Elbadawi, A.H.; Ge, L.; Li, Z.; Liu, S.; Wang, S.; Zhu, Z. Catalytic partial oxidation of methane to syngas: Review of perovskite catalysts and membrane reactors. *Catal. Rev.* **2021**, *63*, 1–67, doi:10.1080/01614940.2020.1743420.
11. Osman, A.I. Catalytic Hydrogen Production from Methane Partial Oxidation: Mechanism and Kinetic Study. *Chem. Eng. Technol.* **2020**, *43*, 641–648, doi:10.1002/ceat.201900339.
12. Choudhary, T.V.; Choudhary, V.R. Energy-Efficient Syngas Production through Catalytic Oxy-Methane Reforming Reactions. *Angew. Chem. Int. Ed.* **2008**, *47*, 1828–1847, doi:10.1002/anie.200701237.
13. Enger, B.C.; Lødeng, R.; Holmen, A. A review of catalytic partial oxidation of methane to synthesis gas with emphasis on reaction mechanisms over transition metal catalysts. *Appl. Catal. A Gen.* **2008**, *346*, 1–27, doi:10.1016/j.apcata.2008.05.018.
14. Campa, M.; Ferraris, G.; Gazzoli, D.; Pettiti, I.; Pietrogiacomi, D. Rhodium supported on tetragonal or monoclinic  $ZrO_2$  as catalyst for the partial oxidation of methane. *Appl. Catal. B Environ.* **2013**, *142–143*, 423–431, doi:10.1016/j.apcatb.2013.05.046.
15. Liu, H.; He, D. Recent Progress on Ni-Based Catalysts in Partial Oxidation of Methane to Syngas. *Catal. Surv. Asia* **2012**, *16*, 53–61, doi:10.1007/s10563-012-9133-4.
16. Jin, R.; Chen, Y.; Li, W.; Cui, W.; Ji, Y.; Yu, C.; Jiang, Y. Mechanism for catalytic partial oxidation of methane to syngas over a Ni/ $Al_2O_3$  catalyst. *Appl. Catal. A Gen.* **2000**, *201*, 71–80, doi:10.1016/s0926-860x(00)00424-5.
17. Dong, W.-S.; Jun, K.-W.; Roh, H.-S.; Liu, Z.-W.; Park, S.-E. Comparative Study on Partial Oxidation of Methane over Ni/ $ZrO_2$ , Ni/ $CeO_2$  and Ni/Ce-ZrO<sub>2</sub> Catalysts. *Catal. Lett.* **2002**, *78*, 215–222, doi:10.1023/a:1014905318290.
18. Yan, Q.G.; Weng, W.Z.; Wan, H.L.; Toghiani, H.; Toghiani, R.K.; Pittman Jr., C.U. Activation of methane to syngas over a Ni/TiO<sub>2</sub> catalyst. *Appl. Catal. A* **2003**, *239*, 43–58.
19. Enger, B.C.; Lødeng, R.; Holmén, A. Evaluation of reactor and catalyst performance in methane partial oxidation over modified nickel catalysts. *Appl. Catal. A Gen.* **2009**, *364*, 15–26, doi:10.1016/j.apcata.2009.05.033.
20. Barbero, J.; Peña, M.A.; Campos-Martin, J.M.; Fierro, J.; Arias, P.L. Support Effect in Supported Ni Catalysts on Their Performance for Methane Partial Oxidation. *Catal. Lett.* **2003**, *87*, 211–218, doi:10.1023/a:1023407609626.
21. Alvarez-Galvan, C.; Melian, M.; Ruiz-Matas, L.; Eslava, J.L.; Navarro, R.M.; Ahmadi, M.; Cuanya, B.R.; Fierro, J.L.G. Partial Oxidation of Methane to Syngas Over Nickel-Based Catalysts: Influence of Support Type, Addition of Rhodium, and Preparation Method. *Front. Chem.* **2019**, *7*, 104, doi:10.3389/fchem.2019.00104.
22. Claridge, J.B.; Green, M.L.H.; Tsang, S.C.; York, A.P.E.; Ashcroft, A.T.; Battle, P.D. A study of carbon deposition on catalysts during the partial oxidation of methane to synthesis gas. *Catal. Lett.* **1993**, *22*, 299–305, doi:10.1007/bf00807237.
23. Mahamulkar, S.; Yin, K.; Agrawal, P.K.; Davis, R.J.; Jones, C.W.; Malek, A.; Shibata, H. Formation and Oxidation/Gasification of Carbonaceous Deposits: A Review. *Ind. Eng. Chem. Res.* **2016**, *55*, 9760–9818, doi:10.1021/acs.iecr.6b02220.
24. Yan, Q.; Wu, T.; Li, J.; Luo, C.; Weng, W.; Yang, L.; Wa, H. Mechanism study of carbon deposition on a Ni/ $Al_2O_3$  catalyst during partial oxidation of methane to syngas. *J. Nat. Gas Chem.* **2000**, *9*, 89–102.
25. Hayek, K.; Kramer, R.; Paál, Z. Metal-support boundary sites in catalysis. *Appl. Catal. A Gen.* **1997**, *162*, 1–15, doi:10.1016/s0926-860x(97)00243-3.
26. Cuanya, B.R. Synthesis and catalytic properties of metal nanoparticles: Size, shape, support, composition, and oxidation state effects. *Thin Solid Films* **2010**, *518*, 3127–3150, doi:10.1016/j.tsf.2010.01.018.
27. Tanabe, K. Surface and catalytic properties of  $ZrO_2$ . *Mater. Chem. Phys.* **1985**, *13*, 347–364, doi:10.1016/0254-0584(85)90064-1.
28. Chuah, G.-K.; Jaenicke, S. The preparation of high surface area zirconia—Influence of precipitating agent and digestion. *Appl. Catal. A Gen.* **1997**, *163*, 261–273, doi:10.1016/s0926-860x(97)00103-8.
29. Jung, K.T.; Bell, A.T. The effects of synthesis and pretreatment conditions on the bulk structure and surface properties of zirconia. *J. Mol. Catal. A Chem.* **2000**, *163*, 27–42, doi:10.1016/s1381-1169(00)00397-6.
30. Hegarty, M.; O’Connor, A.; Ross, J. Syngas production from natural gas using  $ZrO_2$ -supported metals. *Catal. Today* **1998**, *42*, 225–232, doi:10.1016/s0920-5861(98)00096-0.
31. Song, Y.-Q.; He, D.-H.; Xu, B.-Q. Effects of preparation methods of  $ZrO_2$  support on catalytic performances of Ni/ $ZrO_2$  catalysts in methane partial oxidation to syngas. *Appl. Catal. A Gen.* **2008**, *337*, 19–28, doi:10.1016/j.apcata.2007.11.032.
32. Song, Y.-Q.; Liu, H.-M.; He, D.-H. Effects of Hydrothermal Conditions of  $ZrO_2$  on Catalyst Properties and Catalytic Performances of Ni/ $ZrO_2$  in the Partial Oxidation of Methane. *Energy Fuels* **2010**, *24*, 2817–2824, doi:10.1021/ef1000024.
33. Galanov, S.I.; Sidorova, O.I. Effect of a precursor on the phase composition and particle size of the active component of Ni- $ZrO_2$  catalytic systems for the oxidation of methane into syngas. *Russ. J. Phys. Chem. A* **2014**, *88*, 1629–1636.

34. Xu, B.-Q.; Wei, J.-M.; Yu, Y.-T.; Li, Y.; Li, J.-L.; Zhu, Q.-M. Size Limit of Support Particles in an Oxide-Supported Metal Catalyst: Nanocomposite Ni/ZrO<sub>2</sub> for Utilization of Natural Gas. *J. Phys. Chem. B* **2003**, *107*, 5203–5207, doi:10.1021/jp030127l.
35. Pedrero, C.M.; Carrazán, S.G.; Ruiz, P. Preliminary results on the role of the deposition of small amounts of ZrO<sub>2</sub> on Al<sub>2</sub>O<sub>3</sub> support on the partial oxidation of methane and ethane over Rh and Ni supported catalysts. *Catal. Today* **2021**, *363*, 111–121, doi:10.1016/j.cattod.2020.02.010.
36. Martins, R.L.; Schmal, M. Activation of Methane on NiO Nanoparticles with Different Morphologies. *J. Braz. Chem. Soc.* **2014**, *25*, 2399–2408, doi:10.5935/0103-5053.20140266.
37. Yang, J.; Ren, J.; Guo, H.; Qin, X.; Han, B.; Lin, J.; Li, Z. The growth of Ni clusters and their interaction with cubic, monoclinic, and tetragonal ZrO<sub>2</sub> surfaces—a theoretical and experimental study. *RSC Adv.* **2015**, *5*, 59935–59945, doi:10.1039/c5ra07738a.
38. Boudjennad, E.; Chafi, Z.; Ouafek, N.; Ouhenia, S.; Keghouche, N.; Minot, C. Experimental and theoretical study of the Ni-(m-ZrO<sub>2</sub>) interaction. *Surf. Sci.* **2012**, *606*, 1208–1214.
39. Dacquin, J.-P.; Dujardin, C.; Granger, P. Surface reconstruction of supported Pd on LaCoO<sub>3</sub>: Consequences on the catalytic properties in the decomposition of N<sub>2</sub>O. *J. Catal.* **2008**, *253*, 37–49, doi:10.1016/j.jcat.2007.10.023.
40. Scherrer, P. Bestimmung der grössen und der inneren struktur von kolloidteilchen mittels röntgensrahlen. *Nachr. Ges. Wiss. Göttingen* **1918**, *26*, 98–100.
41. Barrett, E.P.; Joyner, L.G.; Halenda, P.P. The Determination of Pore Volume and Area Distributions in Porous Substances. I. Computations from Nitrogen Isotherms. *J. Am. Chem. Soc.* **1951**, *73*, 373–380, doi:10.1021/ja01145a126.
42. Gurvitsch, L. Physicochemical attractive force. *J. Phys. Chem. Soc. Russ.* **1915**, *47*, 805–827.
43. Schneider, C.A.; Rasband, W.S.; Eliceiri, K.W. NIH Image to ImageJ: 25 years of image analysis. *Nat. Methods* **2012**, *9*, 671–675, doi:10.1038/nmeth.2089.
44. Bergeret, G.; Gallezot, P. Particle Size and Dispersion Measurements. In *Handbook of Heterogeneous Catalysis*; Ertl, G., Knözinger, H., Weitkamp, J., Eds.; Wiley-VCH: Weinheim, Germany, 1997; Volume 2, pp. 439–464.
45. Ul-Hamid, A. *A Beginners' Guide to Scanning Electron Microscopy*; Metzler, J.B., Ed.; Springer: Berlin, Germany, 2018; pp. 77–128.
46. Mironova-Ulmane, N.; Kuzmin, A.; Steins, I.; Grabis, J.; Sildos, I.; Pärs, M. Raman scattering in nanosized nickel oxide NiO. *J. Phys. Conf. Ser.* **2007**, *93*, 012039, doi:10.1088/1742-6596/93/1/012039.
47. Wang, W.; Liu, Y.; Xu, C.; Zheng, C.; Wang, G. Synthesis of NiO nanorods by a novel simple precursor thermal decomposition approach. *Chem. Phys. Lett.* **2002**, *362*, 119–122, doi:10.1016/s0009-2614(02)00996-x.
48. Bellido, J.D.A.; Assaf, E.M. Nickel catalysts supported on ZrO<sub>2</sub>, Y<sub>2</sub>O<sub>3</sub>-stabilized ZrO<sub>2</sub> and CaO-stabilized ZrO<sub>2</sub> for the steam reforming of ethanol: Effect of the support and nickel load. *J. Power Sources* **2008**, *177*, 24–32.
49. Zhao, K.; Wang, W.; Li, Z. Highly efficient Ni/ZrO<sub>2</sub> catalysts prepared via combustion method for CO<sub>2</sub> methanation. *J. CO<sub>2</sub> Util.* **2016**, *16*, 236–244, doi:10.1016/j.jcou.2016.07.010.
50. Kesavan, J.K.; Luisetto, I.; Tuti, S.; Meneghini, C.; Iucci, G.; Battocchio, C.; Mobilio, S.; Casciardi, S.; Sisto, R. Nickel supported on YSZ: The effect of Ni particle size on the catalytic activity for CO<sub>2</sub> methanation. *J. CO<sub>2</sub> Util.* **2018**, *23*, 200–211, doi:10.1016/j.jcou.2017.11.015.
51. Mori, H.; Wen, C.-ju; Otomo, J.; Eguchi, K.; Takahashi, H. Investigation of the interaction between NiO and yttria-stabilized zirconia (YSZ) in the NiO/YSZ composite by temperature-programmed reduction technique. *Appl. Catal. A* **2003**, *245*, 79–85.
52. Alderliesten, M. Mean Particle Diameters. From Statistical Definition to Physical Understanding. *J. Biopharm. Stat.* **2005**, *15*, 295–325, doi:10.1081/bip-200048774.
53. Ferrari, A.C.; Robertson, J. Interpretation of Raman spectra of disordered and amorphous carbon. *Phys. Rev. B* **2000**, *61*, 14095–14107, doi:10.1103/physrevb.61.14095.
54. Zhang, X.; Hayward, D.O.; Mingos, D.M.P. Oscillatory behaviour during the partial oxidation of methane over Nickel foils. *Catal. Lett.* **2002**, *83*, 149–155.
55. Zhang, X.; Hayward, D.O.; Mingos, D.M.P. Further Studies on Oscillations over Nickel Wires during the Partial Oxidation of Methane. *Catal. Lett.* **2003**, *86*, 235–243, doi:10.1023/a:1022672219909.
56. Hu, Y.H.; Ruckenstein, E. Catalyst Temperature Oscillations during Partial Oxidation of Methane. *Ind. Eng. Chem. Res.* **1998**, *37*, 2333–2335, doi:10.1021/ie980027f.
57. Wang, M.; Weng, W.; Zheng, H.; Yi, X.; Huang, C.; Wan, H. Oscillations during partial oxidation of methane to synthesis gas over Ru/Al<sub>2</sub>O<sub>3</sub> catalyst. *J. Nat. Gas. Chem.* **2009**, *18*, 300–305, doi:10.1016/s1003-9953(08)60126-7.
58. Zhang, X.; Lee, C.; Hayward, D.; Mingos, D. Oscillatory behaviour observed in the rate of oxidation of methane over metal catalysts. *Catal. Today* **2005**, *105*, 283–294, doi:10.1016/j.cattod.2005.02.040.
59. Bychkov, V.; Tyulerin, Y.; Slinko, M.; Korchak, V. Nonlinear behaviour during methane and ethane oxidation over Ni, Co and Pd catalysts. *Surf. Sci.* **2009**, *603*, 1680–1689, doi:10.1016/j.susc.2008.10.060.
60. Stötzel, J.; Frahm, R.; Kimmerle, B.; Nachtegaal, M.; Grunwaldt, J.-D. Oscillatory behaviour during the catalytic partial oxidation of methane: Following dynamic structural changes of palladium using the QEXAFS technique. *J. Phys. Chem. C* **2012**, *116*, 599–609.
61. Lashina, E.A.; Kaichev, V.V.; Saraev, A.A.; Vinokurov, Z.S.; Chumakova, N.A.; Chumakov, G.A.; Bukhtiyarov, V.I. Experimental Study and Mathematical Modeling of Self-Sustained Kinetic Oscillations in Catalytic Oxidation of Methane over Nickel. *J. Phys. Chem. A* **2017**, *121*, 6874–6886, doi:10.1021/acs.jpca.7b04525.

62. Saraev, A.A.; Vinokurov, Z.S.; Kaichev, V.; Shmakov, A.N.; Bukhtiyarov, V. The origin of self-sustained reaction-rate oscillations in the oxidation of methane over nickel: An operando XRD and mass spectrometry study. *Catal. Sci. Technol.* **2017**, *7*, 1646–1649, doi:10.1039/C6CY02673G.
63. Specchia, S.; Vella, L.D.; Montini, T.; Fornasiero, P. *Hydrogen Production: Prospects and Processes*; Nova Science Publishers Inc.: New York, NY, USA, 2011; pp. 95–139.
64. Pokrovski, K.; Jung, K.T.; Bell, A.T. Investigation of CO and CO<sub>2</sub> Adsorption on Tetragonal and Monoclinic Zirconia. *Langmuir* **2001**, *17*, 4297–4303, doi:10.1021/la001723z.
65. Kouva, S.; Andersin, J.; Honkala, K.; Lehtonen, J.; Lefferts, L.; Kanervo, J. Water and carbon oxides on monoclinic zirconia: Experimental and computational insights. *Phys. Chem. Chem. Phys.* **2014**, *16*, 20650–20664, doi:10.1039/c4cp02742f.
66. Bachiller-Baeza, B.; Rodriguez-Ramos, I.; Guerrero-Ruiz, A. Interaction of Carbon Dioxide with the Surface of Zirconia Poly-morphs. *Langmuir* **1998**, *14*, 3556–3564, doi:10.1021/la970856q.
67. Morterra, C.; Orio, L. Surface characterization of zirconium oxide. II. The interaction with carbon dioxide at ambient temperature. *Mater. Chem. Phys.* **1990**, *24*, 247–268, doi:10.1016/0254-0584(90)90089-s.
68. Bolis, V.; Morterra, C.; Volante, M.; Orio, L.; Fubini, B. Development and suppression of surface acidity on monoclinic zirconia: A spectroscopic and calorimetric investigation. *Langmuir* **1990**, *6*, 695–701, doi:10.1021/la00093a028.
69. Ouyang, F.; Nakayama, A.; Tabada, K.; Suzuki, E. Infrared Study of a Novel Acid-Base Site on ZrO<sub>2</sub> by Adsorbed Probe Molecules. I; Pyridine, Carbon Dioxide, and Formic Acid Adsorption. *J. Phys. Chem. B* **2000**, *104*, 2012–2018.
70. Esconjauregui, S.; Whelan, C.M.; Maex, K. The reasons why metals catalyze the nucleation and growth of carbon nanotubes and other carbon nanomorphologies. *Carbon* **2009**, *47*, 659–669, doi:10.1016/j.carbon.2008.10.047.